

PHYSICS (861)

Aims:

1. To enable candidates to acquire knowledge and to develop an understanding of the terms, facts, concepts, definitions, and fundamental laws, principles and processes in the field of physics.
2. To develop the ability to apply the knowledge and understanding of physics to unfamiliar situations.
3. To develop a scientific attitude through the study of physical sciences.
4. To develop skills in -
 - (a) the practical aspects of handling apparatus, recording observations and
 - (b) Drawing diagrams, graphs, etc.
5. To develop an appreciation of the contribution of physics towards scientific and technological developments and towards human happiness.
6. To develop an interest in the world of physical sciences.

CLASS XI

There will be two papers in the subject:

Paper I: Theory - 3 hours ... 70 marks

Paper II: Practical - 3 hours ... 15 marks

Project Work ... 10 marks

Practical File ... 5 marks

PAPER I- THEORY: 70 Marks

S. NO.	UNIT	TOTAL WEIGHTAGE
1.	Physical World and Measurement	23 Marks
2.	Kinematics	
3.	Laws of Motion	
4.	Work, Energy and Power	17 Marks
5.	Motion of System of Particles and Rigid Body	
6.	Gravitation	
7.	Properties of Bulk Matter	20 Marks
8.	Heat and Thermodynamics	
9.	Behaviour of Perfect Gases and Kinetic Theory of Gases	
10.	Oscillations and Waves	10 Marks
TOTAL		70 Marks

PAPER I -THEORY – 70 MARKS

Note: (i) Unless otherwise specified, only S. I. Units are to be used while teaching and learning, as well as for answering questions.

(ii) All physical quantities to be defined as and when they are introduced along with their units and dimensions.

(iii) Numerical problems are included from all topics except where they are specifically excluded or where only qualitative treatment is required.

1. Physical World and Measurement

Units and Measurements

Measurement: need for measurement; units of measurement; systems of units: fundamental and derived units in SI; measurement of length, mass and time; errors in measurement; significant figures.

Dimensional formulae of physical quantities and constants, dimensional analysis and its applications.

(a) Importance of measurement in scientific studies; physics is a science of measurement. Unit as a reference standard of measurement; essential properties. Systems of units; CGS, FPS, MKS, MKSA, and SI; the seven base units of SI selected by the General Conference of Weights and Measures in 1971 and their definitions, list of fundamental, supplementary and derived physical quantities; their units and symbols (strictly as per rule); subunits and multiple units using prefixes for powers of 10 (from atto for 10^{-18} to tera for 10^{12}); other common units such as fermi, angstrom (now outdated), light year, astronomical unit and parsec. A new unit of mass used in atomic physics is unified atomic mass unit with symbol u (not amu); rules for writing the names of units and their symbols in SI (upper case/lower case.) Derived units (with correct symbols); special names wherever applicable; expression in terms of base units (e.g.: $N = \text{kg m/s}^2$).

(b) Significant figures; their significance; rules for counting the number of significant figures; rules for (a) addition and subtraction, (b) multiplication/division; 'rounding off' the uncertain

digits; order of magnitude as statement of magnitudes in powers of 10; examples from magnitudes of common physical quantities - size, mass, time, etc.

(c) Dimensions of physical quantities; dimensional formula; express derived units in terms of base units ($N = \text{kg m/s}^2$); use symbol [...] for dimensions of or base unit of; e.g.: dimensional formula of force in terms of fundamental quantities written as $[F] = [MLT^{-2}]$. Principle of homogeneity of dimensions. Expressions in terms of SI base units and dimensional formula may be obtained for all physical quantities as and when new physical quantities are introduced.

(d) Use of dimensional analysis to (i) check the dimensional correctness of a formula/ equation; (ii) to obtain the dimensional formula of any derived physical quantity including constants; (iii) to convert units from one system to another; limitations of dimensional analysis.

2. Kinematics

(i) Motion in a Straight Line

Frame of references, Motion in a straight line (one dimension): Position-time graph, speed and velocity.

Elementary concepts of differentiation and integration for describing motion, uniform and non-uniform motion, average speed, velocity, average velocity, instantaneous velocity and uniformly accelerated motion, velocity - time and position - time graphs. Relations for uniformly accelerated motion (graphical treatment).

Frame of reference, concept of point mass, rest and motion; distance and displacement, speed and velocity, average speed and average velocity, uniform velocity, instantaneous speed and instantaneous velocity, acceleration, instantaneous acceleration, $s-t$, $v-t$ and $a-t$ graphs for uniform acceleration and conclusions drawn from these graphs; kinematic equations of motion for objects in uniformly accelerated rectilinear motion derived using graphical,

calculus or analytical method, motion of an object under gravity, (one dimensional motion).

Differentiation as rate of change; examples from physics – speed, acceleration, velocity gradient, etc. Formulae for differentiation of simple functions: x^n , $\sin x$, $\cos x$, e^x and $\ln x$. Simple ideas about integration – mainly. $\int x^n dx$. Both definite and indefinite integrals to be mentioned (elementary calculus not to be evaluated).

(ii) Motion in a Plane

Scalar and Vector quantities with examples. Position and displacement vectors, general vectors and their notations; equality of vectors, addition and subtraction of vectors, Unit vector; resolution of a vector in a plane, rectangular components, Scalar and Vector product of two vectors. Projectile motion and uniform circular motion.

(a) *General Vectors and notation, position and displacement vector. Vectors explained using displacement as a prototype - along a straight line (one dimensional), on a plane surface (two dimensional) and in an open space not confined to a line or a plane (three dimensional); symbol and representation; a scalar quantity, its representation and unit, equality of vectors. Unit vectors denoted by \hat{i} , \hat{j} , \hat{k} orthogonal unit vectors along x, y and z axes respectively. Examples of one dimensional vector $\vec{V}_1 = a\hat{i}$ or $b\hat{j}$ or $c\hat{k}$ where a, b, c are scalar quantities or numbers; $\vec{V}_2 = a\hat{i} + b\hat{j}$ is a two dimensional or planar vector, $\vec{V}_3 = a\hat{i} + b\hat{j} + c\hat{k}$ is a three dimensional or space vector. Concept of null vector and co-planar vectors.*

(b) *Addition: use displacement as an example; obtain triangle law of addition; graphical and analytical treatment; Discuss commutative and associative properties of vector addition (Proof not*

required). Parallelogram Law; sum and difference; derive expressions for magnitude and direction from parallelogram law; special cases; subtraction as special case of addition with direction reversed; use of Triangle Law for subtraction also; if $\vec{a} + \vec{b} = \vec{c}$; $\vec{c} - \vec{a} = \vec{b}$; In a parallelogram, if one diagonal is the sum, the other diagonal is the difference; addition and subtraction with vectors expressed in terms of unit vectors \hat{i} , \hat{j} , \hat{k} ; multiplication of a vector by a real number.

(c) *Use triangle law of addition to express a vector in terms of its components. If $\vec{a} + \vec{b} = \vec{c}$ is an addition fact, $\vec{c} = \vec{a} + \vec{b}$ is a resolution; \vec{a} and \vec{b} are components of \vec{c} . Rectangular components, relation between components, resultant and angle between them. Dot (or scalar) product of vectors $\vec{a} \cdot \vec{b} = ab \cos \theta$; example $W = \vec{F} \cdot \vec{S} = FS \cos \theta$. Special case of $\theta = 0^\circ$, 90° and 180° . Vector (or cross) product $\vec{a} \times \vec{b} = [absin\theta]\hat{n}$; example: torque $\vec{\tau} = \vec{r} \times \vec{F}$; Special cases using unit vectors \hat{i} , \hat{j} , \hat{k} for $\vec{a} \cdot \vec{b}$ and $\vec{a} \times \vec{b}$.*

(d) *Various terms related to projectile motion; obtain equations of trajectory, time of flight, maximum height, horizontal range, instantaneous velocity, [projectile motion on an inclined plane not included]. Examples of projectile motion.*

(e) *Examples of uniform circular motion: details to be covered in unit 3 (d).*

3. Laws of Motion

General concept of force, inertia, Newton's first law of motion; momentum and Newton's second law of motion; impulse; Newton's third law of motion.

Law of conservation of linear momentum and its applications.

Equilibrium of concurrent forces. Friction: Static and kinetic friction, laws of friction, rolling friction, lubrication.

Dynamics of uniform circular motion: Centripetal force, examples of circular motion (vehicle on a level circular road, vehicle on a banked road).

(a) Newton's first law: Statement and explanation; concept of inertia, mass, force; law of inertia; mathematically, if $\sum F=0$, $a=0$.

Newton's second law: $\vec{p}=m\vec{v}$; $\vec{F} \propto \frac{d\vec{p}}{dt}$

$\vec{F}=k \frac{d\vec{p}}{dt}$. Define unit of force so that $k=1$;

$\vec{F} = \frac{d\vec{p}}{dt}$; a vector equation. For classical

physics with v not large and mass m remaining constant, obtain $\vec{F}=m\vec{a}$. For $v \rightarrow c$, m is not constant. Then $m = \frac{m_0}{\sqrt{1-v^2/c^2}}$. Note that $F=ma$ is the

special case for classical mechanics. It is a vector equation. $\vec{a} \parallel \vec{F}$. Also, this can be resolved into three scalar equations $F_x=ma_x$ etc. Application to numerical problems; introduce tension force, normal reaction force. If $a=0$ (body in equilibrium), $F=0$. Statement, derivation and explanation of principle of conservation of linear momentum. Impulse of a force: $F\Delta t = \Delta p$.

Newton's third law. Obtain it using Law of Conservation of linear momentum. Proof of Newton's second law as real law. Systematic solution of problems in mechanics; isolate a part of a system, identify all forces acting on it; draw a free body diagram representing the part as a point and representing all forces by line segments, solve for resultant force which is equal to $m\vec{a}$. Simple problems on "Connected bodies" (not involving two pulleys).

(b) Force diagrams; resultant or net force from Triangle law of Forces, parallelogram law or resolution of forces. Apply net force $\sum \vec{F} = m\vec{a}$. Again for equilibrium $a=0$ and $\sum F=0$. Conditions of equilibrium of a rigid body

under three coplanar forces. Discuss ladder problem.

(c) Friction; classical view and modern view of friction, static friction a self-adjusting force; limiting value; kinetic friction or sliding friction; rolling friction, examples.

Laws of friction: Two laws of static friction; (similar) two laws of kinetic friction; coefficient of friction $\mu_s = f_s(\max)/N$ and $\mu_k = f_k/N$; graphs. Friction as a non-conservative force; motion under friction, net force in Newton's 2nd law is calculated including f_k . Motion along a rough inclined plane – both up and down. Pulling and pushing of a roller. Angle of friction and angle of repose. Lubrication, use of bearings, streamlining, etc.

(d) Angular displacement (θ), angular velocity (ω), angular acceleration (α) and their relations. Concept of centripetal acceleration; obtain an expression for this acceleration using $\Delta \vec{v}$. Magnitude and direction of \vec{a} same as that of $\Delta \vec{v}$; Centripetal acceleration; the cause of this acceleration is a force - also called centripetal force; the name only indicates its direction, it is not a new type of force, motion in a vertical circle; banking of road and railway track (conical pendulum is excluded).

4. Work, Power and Energy

Work done by a constant force and a variable force; kinetic energy, work-energy theorem, power.

Potential energy, potential energy of a spring, conservative forces: conservation of mechanical energy (kinetic and potential energies); Conservative and non-conservative forces. Concept of collision: elastic and inelastic collisions in one and two dimensions.

(i) Work done $W = \vec{F} \cdot \vec{S} = F S \cos \theta$. If F is variable $dW = \vec{F} \cdot d\vec{S}$ and $W = \int dW = \int \vec{F} \cdot d\vec{S}$, for $\vec{F} \parallel d\vec{S}$ $\vec{F} \cdot d\vec{S} = F dS$ therefore, $W = \int F dS$ is the area under the $F-S$ graph or if F can be expressed in terms of S , $\int F dS$ can be evaluated. Example, work done in stretching

a spring $W = \int F dx = \int kx dx = \frac{1}{2} kx^2$. This is also the potential energy stored in the stretched spring $U = \frac{1}{2} kx^2$.

Kinetic energy and its expression, Work-Energy theorem $E=W$. Law of Conservation of Energy; oscillating spring. $U+K = E = K_{max} = U_{max}$ (for $U=0$ and $K=0$ respectively); graph different forms of energy and their transformations. $E = mc^2$ (no derivation). Power $P=W/t$; $P = \vec{F} \cdot \vec{v}$.

(ii) Collision in one dimension; derivation of velocity equation for general case of $m_1 \neq m_2$ and $u_1 \neq u_2 = 0$; Special cases for $m_1 = m_2 = m$; $m_1 > m_2$ or $m_1 < m_2$. Oblique collisions i.e. collision in two dimensions.

5. Motion of System of Particles and Rigid Body

Idea of centre of mass: centre of mass of a two-particle system, momentum conservation and centre of mass motion. Centre of mass of a rigid body; centre of mass of a uniform rod.

Moment of a force, torque, angular momentum, laws of conservation of angular momentum and its applications.

Equilibrium of rigid bodies, rigid body rotation and equations of rotational motion, comparative study of linear and rotational motions.

Moment of inertia, radius of gyration, moments of inertia for simple geometrical objects (no derivation). Statement of parallel and perpendicular axes theorems and their applications.

Definition of centre of mass (cm), centre of mass (cm) for a two particle system $m_1x_1 + m_2x_2 = Mx_{cm}$; differentiating, get the equation for v_{cm} and a_{cm} ; general equation for N particles- many particles system; [need not go into more details]; centre of gravity, principle of moment, discuss ladder problem, concept of a rigid body; kinetic energy of a rigid body rotating about a fixed axis in terms of that of the particles of the body; hence, define moment of inertia and radius of gyration; physical significance of moment of inertia; unit and dimension; depends on mass and axis of rotation; it is rotational inertia; equations of rotational motions. Applications: only expression for the moment of inertia, I (about the symmetry axis) of: (i) a ring; (ii) a solid and a hollow cylinder,

(iii) a thin rod (iv) a solid and a hollow sphere, (v) a disc - only formulae (no derivations required).

(a) Statements of the parallel and perpendicular axes theorems with illustrations [derivation not required]. Simple examples with change of axis.

(b) Definition of torque (vector); $\vec{\tau} = \vec{r} \times \vec{F}$ and angular momentum $\vec{L} = \vec{r} \times \vec{p}$ for a particle (no derivations); differentiate to obtain $d\vec{L}/dt = \vec{\tau}$; similar to Newton's second law of motion (linear); hence $\tau = I \alpha$ and $L = I\omega$; (only scalar equation); Law of conservation of angular momentum; simple applications. Comparison of linear and rotational motions.

6. Gravitation

Kepler's laws of planetary motion, universal law of gravitation. Acceleration due to gravity (g) and its variation with altitude, latitude and depth.

Gravitational potential and gravitational potential energy, escape velocity, orbital velocity of a satellite, Geo-stationary satellites.

(i) Newton's law of universal gravitation; Statement; unit and dimensional formula of universal gravitational constant, G [Cavendish experiment not required]; gravitational acceleration on surface of the earth (g), weight of a body $W = mg$ from $F = ma$.

(ii) Relation between g and G . Derive the expression for variation of g above and below the surface of the earth; graph; mention variation of g with latitude and rotation, (without derivation).

(iii) Gravitational field, intensity of gravitational field and potential at a point in earth's gravitational field. $V_p = W_{ap}/m$. Derive expression (by integration) for the gravitational potential difference $\Delta V = V_B - V_A = G.M(1/r_A - 1/r_B)$; here $V_p = V(r) = -GM/r$; negative sign for attractive force field; define gravitational potential energy of a mass m in the earth's field; expression for gravitational potential

energy $U(r) = W_{ap} = m \cdot V(r) = -G M m/r$; show that $\Delta U = mgh$, for $h \ll R$. Relation between intensity and acceleration due to gravity.

- (iv) Derive expression for the escape velocity of earth using energy consideration; v_e depends on mass of the earth; for moon v_e is less as mass of moon is less; consequence - no atmosphere on the moon.
- (v) Satellites (both natural (moon) and artificial) in uniform circular motion around the earth; Derive the expression for orbital velocity and time period; note the centripetal acceleration is caused (or centripetal force is provided) by the force of gravity exerted by the earth on the satellite; the acceleration of the satellite is the acceleration due to gravity [$g' = g(R/R+h)^2$; $F'_G = mg'$]. Weightlessness; geostationary satellites; conditions for satellite to be geostationary; parking orbit, calculation of its radius and height; basic concept of polar satellites and their uses.
- (vi) Kepler's laws of planetary motion: explain the three laws using diagrams. Proof of third law (for circular orbits only).

7. Properties of Bulk Matter

- (i) Mechanical Properties of Solids: Elastic behaviour of solids, Stress-strain relationship, Hooke's law, Young's modulus, bulk modulus, shear modulus of rigidity, Poisson's ratio; elastic energy (qualitative treatment only).

Elasticity in solids, Hooke's law, Young's modulus and its determination, bulk modulus and shear modulus of rigidity, work done in stretching a wire and strain energy, Poisson's ratio.

- (ii) Mechanical Properties of Fluids

Pressure due to a fluid column; Pascal's law and its applications (hydraulic lift and hydraulic brakes), effect of gravity on fluid pressure.

Viscosity, Stokes' law, terminal velocity, streamline and turbulent flow, critical velocity, Bernoulli's theorem and its applications.

Surface energy and surface tension, angle of contact, excess of pressure across a curved surface, application of surface tension ideas to drops, bubbles and capillary rise.

- (a) Pressure in a fluid, Pascal's Law and its applications, buoyancy (Archimedes Principle).
- (b) General characteristics of fluid flow; equation of continuity $v_1a_1 = v_2a_2$; conditions; applications like use of nozzle at the end of a hose; Bernoulli's principle (theorem); assumptions - incompressible liquid, streamline (steady) flow, non-viscous and irrotational liquid - ideal liquid; derivation of equation; applications of Bernoulli's theorem atomizer, dynamic uplift, Venturimeter, Magnus effect etc.
- (c) Streamline and turbulent flow - examples; streamlines do not intersect (like electric and magnetic lines of force); tubes of flow; number of streamlines per unit area \propto velocity of flow (from equation of continuity $v_1a_1 = v_2a_2$); critical velocity; Reynold's number (significance only) Poiseuille's formula with numericals.
- (d) Viscous drag; Newton's formula for viscosity, co-efficient of viscosity and its units.
Flow of fluids (liquids and gases), laminar flow, internal friction between layers of fluid, between fluid and the solid with which the fluid is in relative motion; examples; viscous drag is a force of friction; mobile and viscous liquids.
- (e) Velocity gradient dv/dx (space rate of change of velocity); viscous drag $F = \eta A dv/dx$; coefficient of viscosity $\eta = F/A (dv/dx)$ depends on the nature of the liquid and its temperature; units: Ns/m^2 and $dyn.s/cm^2 = poise$. $1 poise = 0.1 Ns/m^2$.
Stoke's law, motion of a sphere falling through a fluid, hollow rigid sphere rising to the surface of a liquid, parachute, obtain the expression of

terminal velocity; forces acting; viscous drag, a force proportional to velocity; Stoke's law; v - t graph.

(f) Surface tension (molecular theory), drops and bubbles, angle of contact, work done in stretching a surface and surface energy, capillary rise, measurement of surface tension by capillary (uniform bore) rise method. Excess pressure across a curved surface, application of surface tension for drops and bubbles.

8. Heat and Thermodynamics

(i) Thermal Properties of Matter: Heat, temperature, thermal expansion; thermal expansion of solids, liquids and gases, anomalous expansion of water; specific heat capacity, calorimetry; change of state, specific latent heat capacity.

Heat transfer-conduction, convection and radiation, thermal conductivity, qualitative ideas of Blackbody radiation, Wien's displacement Law and Stefan's law.

(a) Temperature and Heat, measurement of temperature (scales and inter conversion). Ideal gas equation and absolute temperature, thermal expansion in solids, liquids and gases. Specific heat capacity, calorimetry, change of state, latent heat capacity, steady state and temperature gradient. Thermal conductivity; co-efficient of thermal conductivity, Use of good and poor conductors, Searle's experiment, (Lee's Disc method is not required). Convection with examples.

(b) Black body is now called ideal or cavity radiator and black body radiation is cavity radiation; Stefan's law is now known as Stefan Boltzmann law as Boltzmann derived it theoretically. There is multiplicity of technical terms related to thermal radiation - radiant intensity $I(T)$ for total radiant power (energy radiated/second) per unit area of the surface, in W/m^2 , $I(T) = \sigma T^4$; dimension and SI unit of σ . For practical radiators $I = \epsilon \cdot \sigma T^4$ where ϵ (dimension less) is

called emissivity of the surface material; $\epsilon=1$ for ideal radiators. The Spectral radiancy $R(\lambda)$. $I(T) = \int_0^\infty R(\lambda) d\lambda$.

Graph of $R(\lambda)$ vs λ for different temperatures. Area under the graph is $I(T)$. The λ corresponding to maximum value of R is called λ_{max} ; decreases with increase in temperature.

Wien's displacement law; Stefan's law and Newton's law of cooling. [Deductions from Stefan's law not necessary].

(ii) Thermodynamics

Thermal equilibrium and definition of temperature (zeroth law of thermodynamics), heat, work and internal energy. First law of thermodynamics, isothermal and adiabatic processes.

Second law of thermodynamics: reversible and irreversible processes.

(a) Thermal equilibrium and zeroth law of thermodynamics: Self explanatory

(b) First law of thermodynamics.

Concept of heat (Q) as the energy that is transferred (due to temperature difference only) and not stored; the energy that is stored in a body or system as potential and kinetic energy is called internal energy (U). Internal energy is a state property (only elementary ideas) whereas, heat is not; first law is a statement of conservation of energy, when, in general, heat (Q) is transferred to a body (system), internal energy (U) of the system changes and some work W is done by the system; then $Q = \Delta U + W$; also $W = \int pdV$ for working substance - an ideal gas; explain the meaning of symbols (with examples) and sign convention carefully (as used in physics: $Q > 0$ when added to a system, $\Delta U > 0$ when U increases or temperature rises, and $W > 0$ when work is done by the system). Special cases for $Q = 0$ (adiabatic), $\Delta U = 0$ (isothermal) and $W = 0$ (isochoric).

(c) Isothermal and adiabatic changes in a perfect gas described in terms of PV graphs; $PV = \text{constant}$ (Isothermal) and $PV' = \text{constant}$ (adiabatic); joule and calorie relation (derivation of $PV' = \text{constant}$ not required).

Note that $1 \text{ cal} = 4.186 \text{ J}$ exactly and J (so-called mechanical equivalent of heat) should not be used in equations. In equations, it is understood that each term as well as the LHS and RHS are in the same units; it could be all joules or all calories.

(d) Derive an expression for work done in isothermal and adiabatic processes; principal and molar heat capacities; C_p and C_v ; relation between C_p and C_v ($C_p - C_v = R$). Work done as area bounded by PV graph.

(e) Second law of thermodynamics, Carnot's cycle. Some practical applications.

Only one statement each in terms of Kelvin's impossible steam engine and Clausius' impossible refrigerator. Brief explanation of the law. Reversible and irreversible processes, Heat engine; Carnot's cycle - describe realisation from source and sink of infinite thermal capacity, thermal insulation, etc. Explain using pV graph (isothermal process and adiabatic process) expression and numericals (without derivation) for efficiency $\eta = 1 - T_2/T_1$.

9. Behaviour of Perfect Gases and Kinetic Theory of Gases

(i) Kinetic Theory: Equation of state of a perfect gas, work done in compressing a gas. Kinetic theory of gases - assumptions, concept of pressure. Kinetic interpretation of temperature; rms speed of gas molecules; degrees of freedom, law of equi-partition of energy (statement only) and application to specific heat capacities of gases; concept of mean free path, Avogadro's number.

(a) Kinetic Theory of gases; derive $p = 1/3 \rho \bar{c}^2$ from the assumptions and applying Newton's laws of motion. The average thermal velocity (rms value) $c_{\text{rms}} = \sqrt{3p/\rho}$; calculations for air, hydrogen and their

comparison with common speeds. Effect of temperature and pressure on rms speed of gas molecules.

[Note that $pV = nRT$ the ideal gas equation cannot be derived from kinetic theory of ideal gas. Hence, neither can other gas laws; $pV = nRT$ is an experimental result. Comparing this with $p = \frac{1}{3} \rho \bar{c}^2$, from kinetic theory of gases, a kinetic interpretation of temperature can be obtained as explained in the next subunit].

(b) From kinetic theory for an ideal gas (obeying all the assumptions especially no intermolecular attraction and negligibly small size of molecules, we get $p = (1/3)\rho \bar{c}^2$ or $pV = (1/3)M \bar{c}^2$. (No further, as temperature is not a concept of kinetic theory). From experimentally obtained gas laws, we have the ideal gas equation (obeyed by some gases at low pressure and high temperature) $pV = RT$ for one mole. Combining these two results (assuming they can be combined), $RT = (1/3)M \bar{c}^2 = (2/3) \cdot \frac{1}{2} M \bar{c}^2 = (2/3)K$; Hence, kinetic energy of 1 mole of an ideal gas $K = (3/2)RT$. Average K for 1 molecule $= K/N = (3/2) RT/N = (3/2) kT$ where k is Boltzmann's constant. So, temperature T can be interpreted as a measure of the average kinetic energy of the molecules of a gas.

(c) Degrees of freedom and calculation of specific heat capacities for all types of gases. Concept of the law of equipartition of energy (derivation not required). Concept of mean free path and Avogadro's number N_A .

10. Oscillations and Waves

(i) Oscillations: Periodic motion, time period, frequency, displacement as a function of time, periodic functions. Simple harmonic motion (S.H.M) and its equation; phase; oscillations of a spring, restoring force and force constant; energy in S.H.M., Kinetic and potential energies; simple pendulum and derivation of expression for its time period.

Simple harmonic motion. Periodic motion, time period T and frequency f , $f=1/T$; uniform circular motion and its projection on a diameter defines SHM; displacement, amplitude, phase and epoch, velocity, acceleration, time period; characteristics of SHM; Relation between linear simple harmonic motion and uniform circular motion. Differential equation of SHM, $d^2y/dt^2 + \omega^2y = 0$ from the nature of force acting $F = -k y$; solution $y = A \sin(\omega t + \phi_0)$ where $\omega^2 = k/m$; obtain expressions for velocity, acceleration, time period T and frequency f . Graphical representation of displacement, velocity and acceleration. Examples, simple pendulum, a mass m attached to a spring of spring constant k . Derivation of time period of simple harmonic motion of a simple pendulum, mass on a spring (horizontal and vertical oscillations) Kinetic and potential energy at a point in simple harmonic motion. Total energy $E = U+K$ (potential +kinetic) is conserved. Draw graphs of U , K and E Verses y .

(ii) Waves: Wave motion, Transverse and longitudinal waves, speed of wave motion, displacement relation for a progressive wave, principle of superposition of waves, reflection of waves, standing waves in strings and organ pipes, fundamental mode and harmonics, Beats.

(a) *Transverse and longitudinal waves; characteristics of a harmonic wave; graphical representation of a harmonic wave. Distinction between transverse and longitudinal waves; examples; displacement, amplitude, time period, frequency, wavelength, derive $v = f\lambda$; graph of displacement with time/position,*

label time period/wavelength and amplitude, equation of a progressive harmonic (sinusoidal) wave, $y = A \sin(kx \pm \omega t)$ where k is a propagation factor and equivalent equations.

(b) *Production and propagation of sound as a wave motion; mechanical wave requires a medium; general formula for speed of sound (no derivation). Newton's formula for speed of sound in air; experimental value; Laplace's correction; variation of speed v with changes in pressure, density, humidity and temperature. Speed of sound in liquids and solids - brief introduction only. Concept of supersonic and ultrasonic waves.*

(c) *Principle of superposition of waves; interference (simple ideas only); dependence of combined wave form, on the relative phase of the interfering waves; qualitative only - illustrate with wave representations. Beats (qualitative explanation only); number of beats produced per second = difference in the frequencies of the interfering waves. Standing waves or stationary waves; formation by two identical progressive waves travelling in opposite directions (e.g., along a string, in an air column - incident and reflected waves); obtain $y = y_1 + y_2 = [2 y_m \sin(kx)] \cos(\omega t)$ using equations of the travelling waves; variation of the amplitude $A = 2 y_m \sin(kx)$ with location (x) of the particle; nodes and antinodes; compare standing waves with progressive waves.*

(d) *Laws of vibrations of a stretched string. Obtain equation for fundamental frequency $f_0 = (\gamma l) \sqrt{T/m}$; sonometer.*

(e) *Modes of vibration of strings and air columns (closed and open pipes); standing waves with nodes and antinodes; also in resonance with the periodic force exerted usually by a tuning fork; sketches of various modes of vibration; obtain expressions for fundamental frequency and various harmonics and overtones; mutual relations.*

PAPER II

PRACTICAL WORK- 15 Marks

Given below is a list of required experiments. Teachers may add to this list, keeping in mind the general pattern of questions asked in the annual examinations.

In each experiment, students are expected to record their observations in a tabular form with units at the column head. Students should plot an appropriate graph, work out the necessary calculations and arrive at the result.

Students are required to have completed all experiments from the given list (excluding demonstration experiments):

1. To measure the diameter of a spherical body using Vernier calipers. Calculate its volume with appropriate significant figures. Also measure its volume using a graduated cylinder and compare the two.
2. Find the diameter of a wire using a micrometer screw gauge and determine percentage error in cross sectional area.
3. Determine radius of curvature of a spherical surface like watch glass by a spherometer.
4. Equilibrium of three concurrent coplanar forces. To verify the parallelogram law of forces and to determine weight of a body.
5. (i) Inclined plane: To find the downward force acting along the inclined plane on a roller due to gravitational pull of earth and to study its relationship with angle of inclination by plotting graph between force and $\sin \theta$.
(ii) Friction: To find the force of limiting friction for a wooden block placed on horizontal surface and to study its relationship with normal reaction. To determine the coefficient of friction.
6. To find the acceleration due to gravity by measuring the variation in time period (T) with effective length (L) of a simple pendulum; plot graphs of T vs \sqrt{L} and T^2 vs L. Determine effective length of the seconds pendulum from T^2 vs L graph.
7. To find the force constant of a spring and to study variation in time period of oscillation with mass m of a body suspended by the spring. To

find acceleration due to gravity by plotting a graph of T against \sqrt{m} .

8. Boyle's Law: To study the variation in volume with pressure for a sample of air at constant temperature by plotting graphs between p and $1/V$ and between p and V.
9. Cooling curve: To study the fall in temperature of a body (like hot water or liquid in calorimeter) with time. Find the slope of the curve at four different temperatures of the hot body and hence, deduce Newton's law of cooling.
10. To study the variation in frequency of air column with length using resonance column apparatus or a long cylindrical vessel and a set of tuning forks. Hence, determine velocity of sound in air at room temperature.
11. To determine frequency of a tuning fork using a sonometer.
12. To determine specific heat capacity of a solid using a calorimeter.

Demonstration Experiments (The following experiments are to be demonstrated by the teacher):

1. Searle's method to determine Young modulus of elasticity.
2. Capillary rise method to determine surface tension of water.
3. Determination of coefficient of viscosity of a given viscous liquid by terminal velocity method.

PROJECT WORK AND PRACTICAL FILE – 15 Marks

Project Work – 10 Marks

All candidates will be required to do **one** project involving some Physics related topic/s, under the guidance and regular supervision of the Physics teacher. Candidates are to prepare a technical report including an abstract, some theoretical discussion, experimental setup, observations with tables of data collected, analysis and discussion of results, deductions, conclusion, etc. (after the draft has been approved by the teacher). The report should be kept simple, but neat and elegant. Teachers may assign or students may choose any one project of their choice.

Suggested Evaluation criteria:

▪ Title and Abstract (summary)
▪ Introduction / purpose
▪ Contents/Presentation
▪ Analysis/ material aid (graph, data, structure, pie charts, histograms, diagrams, etc.)
▪ Originality of work
▪ Conclusion/comments

Practical File – 5 Marks

Teachers are required to assess students on the basis of the Physics practical file maintained by them during the academic year.

NOTE: For guidelines regarding Project Work, please refer to Class XII.

CLASS XII

There will be two papers in the subject:

Paper I: Theory - 3 hours ... 70 marks

Paper II: Practical - 3 hours ... 15 marks

Project Work ... 10 marks

Practical File ... 5 marks

PAPER I- THEORY: 70 Marks

S. NO.	UNIT	TOTAL WEIGHTAGE
1.	Electrostatics	14 Marks
2.	Current Electricity	
3.	Magnetic Effects of Current and Magnetism	16 Marks
4.	Electromagnetic Induction and Alternating Currents	
5.	Electromagnetic Waves	2 Marks
6.	Optics	18 Marks
7.	Dual Nature of Radiation and Matter	7 Marks
8.	Atoms and Nuclei	6 Marks
9.	Electronic Devices	7 Marks
TOTAL		70 Marks

PAPER I -THEORY- 70 Marks

Note: (i) Unless otherwise specified, only S. I. Units are to be used while teaching and learning, as well as for answering questions.

(ii) All physical quantities to be defined as and when they are introduced along with their units and dimensions.

(iii) Numerical problems are included from all topics except where they are specifically excluded or where only qualitative treatment is required.

1. Electrostatics

(i) Electric Charges and Fields

Electric charges; conservation and quantisation of charge, Coulomb's law; superposition principle and continuous charge distribution.

Electric field: electric field due to a point charge, electric field lines, electric dipole, electric field due to a dipole, torque on a dipole in uniform electric field.

Electric flux, Gauss's theorem in Electrostatics and its applications to find field due to infinitely long straight wire, uniformly charged infinite plane sheet and uniformly charged thin spherical shell.

(a) Coulomb's law, S.I. unit of charge; permittivity of free space and of dielectric medium. Frictional electricity, electric charges (two types); repulsion and attraction; simple atomic structure - electrons and ions; conductors and insulators; quantization and conservation of electric charge; Coulomb's law in vector form; (position coordinates r_1, r_2 not necessary). Comparison with Newton's law of gravitation; Superposition principle $(\vec{F}_1 = \vec{F}_{12} + \vec{F}_{13} + \vec{F}_{14} + \dots)$.

(b) Concept of electric field and its intensity; examples of different fields; gravitational, electric and magnetic; Electric field due to a point charge $\vec{E} = \vec{F} / q_0$ (q_0 is a test charge); \vec{E} for a group of charges (superposition principle); a point charge q in an electric

field \vec{E} experiences an electric force $\vec{F}_E = q\vec{E}$. Intensity due to a continuous distribution of charge i.e. linear, surface and volume.

(c) Electric lines of force: A convenient way to visualize the electric field; properties of lines of force; examples of the lines of force due to (i) an isolated point charge (+ve and - ve); (ii) dipole, (iii) two similar charges at a small distance; (iv) uniform field between two oppositely charged parallel plates.

(d) Electric dipole and dipole moment; derivation of the \vec{E} at a point, (1) on the axis (end on position) (2) on the perpendicular bisector (equatorial i.e. broad side on position) of a dipole, also for $r \gg 2l$ (short dipole); dipole in a uniform electric field; net force zero, torque on an electric dipole: $\vec{\tau} = \vec{p} \times \vec{E}$ and its derivation.

(e) Gauss' theorem: the flux of a vector field; $Q = vA$ for velocity vector $\vec{v} \parallel \vec{A}$, \vec{A} is area vector. Similarly, for electric field \vec{E} , electric flux $\phi_E = EA$ for $\vec{E} \parallel \vec{A}$ and $\phi_E = \vec{E} \cdot \vec{A}$ for uniform \vec{E} . For non-uniform field $\phi_E = \int d\phi = \int \vec{E} \cdot d\vec{A}$. Special cases for $\theta = 0^\circ, 90^\circ$ and 180° . Gauss' theorem, statement: $\phi_E = q/\epsilon_0$ or $\phi_E = \oint \vec{E} \cdot d\vec{A} = \frac{q}{\epsilon_0}$ where ϕ_E is for a closed surface; q is the net charge enclosed, ϵ_0 is the permittivity of free space. Essential properties of a Gaussian surface.

Applications: Obtain expression for \vec{E} due to 1. an infinite line of charge, 2. a uniformly charged infinite plane thin sheet, 3. a thin hollow spherical shell (inside, on the surface and outside). Graphical variation of E vs r for a thin spherical shell.

(ii) Electrostatic Potential, Potential Energy and Capacitance

Electric potential, potential difference, electric potential due to a point charge, a dipole and system of charges; equipotential surfaces, electrical potential energy of a system of two point charges and of electric dipole in an electrostatic field.

Conductors and insulators, free charges and bound charges inside a conductor. Dielectrics and electric polarisation, capacitors and capacitance, combination of capacitors in series and in parallel. Capacitance of a parallel plate capacitor, energy stored in a capacitor (No derivation, formulae only).

(a) *Concept of potential, potential difference and potential energy. Equipotential surface and its properties. Obtain an expression for electric potential at a point due to a point charge; graphical variation of E and V vs r , $V_p = W/q_0$; hence $V_A - V_B = W_{BA}/q_0$ (taking q_0 from B to A) = $(q/4\pi\epsilon_0)(1/r_A - 1/r_B)$; derive this equation; also $V_A = q/4\pi\epsilon_0 \cdot 1/r_A$; for $q > 0$, $V_A > 0$ and for $q < 0$, $V_A < 0$. For a collection of charges V = algebraic sum of the potentials due to each charge; potential due to a dipole on its axial line and equatorial line; also at any point for $r \gg 2l$ (short dipole). Potential energy of a point charge (q) in an electric field \vec{E} , placed at a point P where potential is V , is given by $U = qV$ and $\Delta U = q(V_A - V_B)$. The electrostatic potential energy of a system of two charges = work done $W_{21} = W_{12}$ in assembling the system; U_{12} or $U_{21} = (1/4\pi\epsilon_0) q_1 q_2 / r_{12}$. For a system of 3 charges $U_{123} = U_{12} + U_{13} + U_{23}$*

$= \frac{1}{4\pi\epsilon_0} \left(\frac{q_1 q_2}{r_{12}} + \frac{q_1 q_3}{r_{13}} + \frac{q_2 q_3}{r_{23}} \right)$. For a dipole in a uniform electric field, derive an expression of the electric potential energy $U_E = -\vec{p} \cdot \vec{E}$, special cases for $\phi = 0^\circ, 90^\circ$ and 180° .

(b) *Capacitance of a conductor $C = Q/V$; obtain the capacitance of a parallel-plate capacitor ($C = \epsilon_0 A/d$) and equivalent*

capacitance for capacitors in series and parallel combinations. Expression for energy stored ($U = \frac{1}{2} CV^2$

$$= \frac{1}{2} QV = \frac{1}{2} \frac{Q^2}{C}) \text{ and energy density.}$$

(c) *Dielectric constant $K = C'/C$; this is also called relative permittivity $K = \epsilon_r = \epsilon/\epsilon_0$; elementary ideas of polarization of matter in a uniform electric field qualitative discussion; induced surface charges weaken the original field; results in reduction in \vec{E} and hence, in pd , (V); for charge remaining the same $Q = CV = C'V' = K \cdot CV$; $V' = V/K$; and $E' = \frac{E}{K}$; if the Capacitor is kept connected with the source of emf, V is kept constant $V = Q/C = Q'/C'$; $Q' = C'V = K \cdot CV = K \cdot Q$ increases; For a parallel plate capacitor with a dielectric in between, $C' = KC = K \cdot \epsilon_0$. $A/d = \epsilon_r \cdot \epsilon_0 \cdot A/d$. Then $C' = \frac{\epsilon_0 A}{d/\epsilon_r}$; for a capacitor partially filled dielectric, capacitance, $C' = \epsilon_0 A/(d-t+t/\epsilon_r)$.*

2. Current Electricity

Mechanism of flow of current in conductors. Mobility, drift velocity and its relation with electric current; Ohm's law and its proof, resistance and resistivity and their relation to drift velocity of electrons; V-I characteristics (linear and non-linear), electrical energy and power, electrical resistivity and conductivity. Temperature dependence of resistance and resistivity.

Internal resistance of a cell, potential difference and emf of a cell, combination of cells in series and in parallel, Kirchhoff's laws and simple applications, Wheatstone bridge, metre bridge. Potentiometer - principle and its applications to measure potential difference, to compare emf of two cells; to measure internal resistance of a cell.

(a) Free electron theory of conduction; acceleration of free electrons, relaxation time τ ; electric current $I = Q/t$; concept of drift velocity and electron mobility. Ohm's law, current density $J = I/A$; experimental verification, graphs and slope, ohmic and non-ohmic conductors; obtain the relation $I = v_{d}enA$. Derive $\sigma = ne^2\tau/m$ and $\rho = m/ne^2\tau$; effect of temperature on resistivity and resistance of conductors and semiconductors and graphs. Resistance $R = V/I$; resistivity ρ , given by $R = \rho l/A$; conductivity and conductance; Ohm's law as $\vec{J} = \sigma \vec{E}$.

(b) Electrical energy consumed in time t is $E = Pt = VIt$; using Ohm's law $E = \left(V^2/R\right)t = I^2Rt$. Potential difference $V = P/I$; $P = VI$; Electric power consumed $P = VI = V^2/R = I^2R$; commercial units; electricity consumption and billing.

(c) The source of energy of a seat of emf (such as a cell) may be electrical, mechanical, thermal or radiant energy. The emf of a source is defined as the work done per unit charge to force them to go to the higher point of potential (from -ve terminal to +ve terminal inside the cell) so, $\varepsilon = dW/dq$; but $dq = Idt$; $dW = \varepsilon dq = \varepsilon Idt$. Equating total work done to the work done across the external resistor R plus the work done across the internal resistance r ; $\varepsilon Idt = I^2R dt + I^2r dt$; $\varepsilon = I(R + r)$; $I = \varepsilon/(R + r)$; also $IR + Ir = \varepsilon$ or $V = \varepsilon - Ir$ where Ir is called the back emf as it acts against the emf ε ; V is the terminal pd. Derivation of formulae for combination for identical cells in series, parallel and mixed grouping. Parallel combination of two cells of unequal emf. Series combination of n cells of unequal emf.

(d) Statement and explanation of Kirchhoff's laws with simple examples. The first is a conservation law for charge and the 2nd is law of conservation of energy. Note change in potential across a resistor $\Delta V = IR < 0$ when we go 'down' with the current (compare with flow of water down a river), and $\Delta V = IR > 0$ if we go up against the current across the resistor. When we go through a cell, the -ve

terminal is at a lower level and the +ve terminal at a higher level, so going from -ve to +ve through the cell, we are going up and $\Delta V = +\varepsilon$ and going from +ve to -ve terminal through the cell, we are going down, so $\Delta V = -\varepsilon$. Application to simple circuits. Wheatstone bridge; right in the beginning take $I_g = 0$ as we consider a balanced bridge, derivation of $R_1/R_2 = R_3/R_4$ [Kirchhoff's law not necessary]. Metre bridge is a modified form of Wheatstone bridge, its use to measure unknown resistance. Here $R_3 = l_1\rho$ and $R_4 = l_2\rho$; $R_3/R_4 = l_1/l_2$. Principle of Potentiometer: fall in potential $\Delta V \propto \Delta l$; auxiliary emf ε_1 is balanced against the fall in potential V_1 across length l_1 . $\varepsilon_1 = V_1 = Kl_1$; $\varepsilon_1/\varepsilon_2 = l_1/l_2$; potentiometer as a voltmeter. Potential gradient and sensitivity of potentiometer. Use of potentiometer: to compare emfs of two cells, to determine internal resistance of a cell.

3. Magnetic Effects of Current and Magnetism

(i) Moving charges and magnetism

Concept of magnetic field, Oersted's experiment. Biot - Savart law and its application. Ampere's Circuital law and its applications to infinitely long straight wire, straight solenoids (only qualitative treatment). Force on a moving charge in uniform magnetic and electric fields. Force on a current-carrying conductor in a uniform magnetic field, force between two parallel current-carrying conductors-definition of ampere, torque experienced by a current loop in uniform magnetic field; moving coil galvanometer - its sensitivity. Conversion of galvanometer into an ammeter and a voltmeter.

(ii) Magnetism and Matter

A current loop as a magnetic dipole, its magnetic dipole moment, magnetic dipole moment of a revolving electron, magnetic field intensity due to a magnetic dipole (bar magnet) on the axial line and equatorial line (Qualitative only) torque on a magnetic dipole (bar magnet) in a uniform magnetic field; bar magnet as an equivalent solenoid. Diamagnetic, paramagnetic, and

ferromagnetic substances, with examples. Electromagnets and factors affecting their strengths, permanent magnets.

- (a) Only historical introduction through Oersted's experiment. [Ampere's swimming rule not included]. Biot-Savart law and its vector form; application; derive the expression for B (i) at the centre of a circular loop carrying current; (ii) at any point on its axis. Current carrying loop as a magnetic dipole. Ampere's Circuital law: statement and brief explanation. Apply it to obtain \vec{B} near a long wire carrying current and for a solenoid. Only formula of \vec{B} due to a finitely long conductor.
- (b) Force on a moving charged particle in magnetic field $\vec{F}_B = q(\vec{v} \times \vec{B})$; special cases, modify this equation substituting $d\vec{l}/dt$ for v and I for q/dt to yield $\vec{F} = I\vec{dl} \times \vec{B}$ for the force acting on a current carrying conductor placed in a magnetic field. Derive the expression for force between two long and parallel wires carrying current, hence, define ampere (the base SI unit of current) and hence, coulomb; from $Q = It$. Lorentz force.
- (c) Derive the expression for torque on a current carrying loop placed in a uniform \vec{B} , using $\vec{F} = I\vec{l} \times \vec{B}$ and $\vec{\tau} = \vec{r} \times \vec{F}$; $\tau = NIAB \sin \phi$ for N turns $\vec{\tau} = \vec{m} \times \vec{B}$, where the dipole moment $\vec{m} = NI\vec{A}$, unit: $A.m^2$. A current carrying loop is a magnetic dipole; directions of current and \vec{B} and \vec{m} using right hand rule only; no other rule necessary. Mention orbital magnetic moment of an electron in Bohr model of H atom. Concept of radial magnetic field. Moving coil galvanometer; construction, principle, working, theory $I = k\phi$, current and voltage sensitivity. Shunt. Conversion of galvanometer into ammeter and voltmeter of given range.

(d) Magnetic field represented by the symbol \vec{B} is now defined by the equation $\vec{F} = q_o(\vec{v} \times \vec{B})$; \vec{B} is not to be defined in terms of force acting on a unit pole, etc.; note the distinction of \vec{B} from \vec{E} is that \vec{B} forms closed loops as there are no magnetic monopoles, whereas \vec{E} lines start from +ve charge and end on -ve charge. Magnetic field lines due to a magnetic dipole (bar magnet). Magnetic field in end-on and broadside-on positions (No derivations). Magnetic flux $\phi = \vec{B} \cdot \vec{A} = BA$ for B uniform and $\vec{B} \parallel \vec{A}$; i.e. area held perpendicular to \vec{B} . For $\phi = BA(\vec{B} \parallel \vec{A})$, $B = \phi/A$ is the flux density [SI unit of flux is weber (Wb)]; but note that this is not correct as a defining equation as \vec{B} is vector and ϕ and ϕ/A are scalars, unit of B is tesla (T) equal to 10^4 gauss. For non-uniform \vec{B} field, $\phi = \int d\phi = \int \vec{B} \cdot d\vec{A}$.

- (e) Properties of diamagnetic, paramagnetic and ferromagnetic substances; their susceptibility and relative permeability.

It is better to explain the main distinction, the cause of magnetization (M) is due to magnetic dipole moment (m) of atoms, ions or molecules being 0 for dia, > 0 but very small for para and > 0 and large for ferromagnetic materials; few examples; placed in external \vec{B} , very small (induced) magnetization in a direction opposite to \vec{B} in dia, small magnetization parallel to \vec{B} for para, and large magnetization parallel to \vec{B} for ferromagnetic materials; this leads to lines of \vec{B} becoming less dense, more dense and much more dense in dia, para and ferro, respectively; hence, a weak repulsion for dia, weak attraction for para and strong attraction for ferro magnetic material. Also, a small bar suspended in the horizontal plane becomes perpendicular to the \vec{B} field for dia and parallel to \vec{B}

for para and ferro. Defining equation $H = (B/\mu_0) - M$; the magnetic properties, susceptibility $\chi_m = (M/H) < 0$ for dia (as M is opposite H) and >0 for para, both very small, but very large for ferro; hence relative permeability $\mu_r = (1 + \chi_m) < 1$ for dia, > 1 for para and $>>1$ (very large) for ferro; further, $\chi_m \propto 1/T$ (Curie's law) for para, independent of temperature (T) for dia and depends on T in a complicated manner for ferro; on heating ferro becomes para at Curie temperature. Electromagnet: its definition, properties and factors affecting the strength of electromagnet; selection of magnetic material for temporary and permanent magnets and core of the transformer on the basis of retentivity and coercive force [B-H loop and its significance, retentivity and coercive force (Qualitative only)].

4. Electromagnetic Induction and Alternating Currents

(i) Electromagnetic Induction

Faraday's laws, induced emf and current; Lenz's Law, eddy currents. Self-induction and mutual induction. Transformer.

(ii) Alternating Current

Peak value, mean value and RMS value of alternating current/voltage; their relation in sinusoidal case; reactance and impedance; LC oscillations (qualitative treatment only), LCR series circuit, resonance; power in AC circuits, wattless current. AC generator.

(a) Electromagnetic induction, Magnetic flux, change in flux, rate of change of flux and induced emf; Faraday's laws. Lenz's law, conservation of energy; motional emf $\varepsilon = Blv$, and power $P = (Blv)^2/R$; eddy currents (qualitative);

(b) Self-Induction, coefficient of self-inductance, $\phi = LI$ and $L = \varepsilon/dI/dt$; henry = volt. Second/ampere, expression for coefficient of self-inductance of a solenoid $L = \frac{\mu_0 N^2 A}{l} = \mu_0 n^2 A \times l$.

Mutual induction and mutual inductance (M), flux linked $\phi_2 = MI_1$; induced emf $\varepsilon_2 = \frac{d\phi_2}{dt} = M \frac{dI_1}{dt}$. Definition of M as

$$M = \frac{\varepsilon_2}{dI_1} \text{ or } M = \frac{\phi_2}{I_1} \text{. SI unit}$$

henry. Expression for coefficient of mutual inductance of two coaxial solenoids.

$M = \frac{\mu_0 N_1 N_2 A}{l} = \mu_0 n_1 N_2 A$ Induced emf opposes changes, back emf is set up, eddy currents.

Transformer (ideal coupling): principle, working and uses; step up and step down; efficiency and applications including transmission of power, energy losses and their minimisation.

(c) Sinusoidal variation of V and I with time, for the output from an ac generator; time period, frequency and phase changes; obtain mean values of current and voltage, obtain relation between RMS value of V and I with peak values in sinusoidal cases only.

(d) Variation of voltage and current in a.c. circuits consisting of only a resistor, only an inductor and only a capacitor (phasor representation), phase lag and phase lead. May apply Kirchhoff's law and obtain simple differential equation (SHM type), $V = V_0 \sin \omega t$, solution $I = I_0 \sin \omega t$, $I_0 \sin(\omega t + \pi/2)$ and $I_0 \sin(\omega t - \pi/2)$ for pure R , C and L circuits respectively. Draw phase (or phasor) diagrams showing voltage and current and phase lag or lead, also showing resistance R , inductive reactance X_L ; ($X_L = \omega L$) and capacitive reactance X_C ; ($X_C = 1/\omega C$). Graph of X_L and X_C vs f .

(e) The LCR series circuit: Use phasor diagram method to obtain expression for I and V , the pd across R , L and C ; and the net phase lag/lead; use the results of 4(e), V lags I by $\pi/2$ in a capacitor, V leads I by $\pi/2$ in an inductor, V and I are

in phase in a resistor, I is the same in all three; hence draw phase diagram, combine V_L and V_C (in opposite phase; phasors add like vectors) to give $V = V_R + V_L + V_C$ (phasor addition) and the max. values are related by $V_m^2 = V_{Rm}^2 + (V_{Lm} - V_{Cm})^2$ when $V_L > V_C$. Substituting $pd = \text{current} \times \text{resistance}$ or reactance , we get $Z^2 = R^2 + (X_L - X_C)^2$ and $\tan \phi = (V_{Lm} - V_{Cm})/V_{Rm} = (X_L - X_C)/R$ giving $I = I_m \sin(\omega t - \phi)$ where $I_m = V_m/Z$ etc. Special cases for RL and RC circuits. [May use Kirchoff's law and obtain the differential equation] Graph of Z vs f and I vs f .

(f) Power P associated with LCR circuit = $1/2 V_o I_o \cos \phi = V_{rms} I_{rms} \cos \phi = I_{rms}^2 R$; power absorbed and power dissipated; electrical resonance; bandwidth of signals and Q factor (no derivation); oscillations in an LC circuit ($\omega_0 = 1/\sqrt{LC}$). Average power consumed averaged over a full cycle $\bar{P} = (1/2) V_o I_o \cos \phi$. Power factor $\cos \phi = R/Z$. Special case for pure R , L and C ; choke coil (analytical only), X_L controls current but $\cos \phi = 0$, hence $\bar{P} = 0$, wattless current; LC circuit; at resonance with $X_L = X_C$, $Z = Z_{min} = R$, power delivered to circuit by the source is maximum, resonant frequency

$$f_0 = \frac{1}{2\pi\sqrt{LC}}.$$

(g) Simple a.c. generators: Principle, description, theory, working and use. Variation in current and voltage with time for a.c. and d.c. Basic differences between a.c. and d.c.

5. Electromagnetic Waves

Basic idea of displacement current. Electromagnetic waves, their characteristics, their transverse nature (qualitative ideas only). Complete electromagnetic spectrum starting from radio waves to gamma rays: elementary facts of electromagnetic waves and their uses.

Concept of displacement current, qualitative descriptions only of electromagnetic spectrum; common features of all regions of electromagnetic spectrum including transverse nature (E and B perpendicular to \vec{c}); special features of the common classification (gamma rays, X rays, UV rays, visible light, IR, microwaves, radio and TV waves) in their production (source), detection and other properties; uses; approximate range of λ or f or at least proper order of increasing f or λ .

6. Optics

(i) Ray Optics and Optical Instruments

Ray Optics: Reflection of light by spherical mirrors, mirror formula, refraction of light at plane surfaces, total internal reflection and its applications, optical fibres, refraction at spherical surfaces, lenses, thin lens formula, lens maker's formula, magnification, power of a lens, combination of thin lenses in contact, combination of a lens and a mirror, refraction and dispersion of light through a prism.

Optical instruments: Microscopes and astronomical telescopes (reflecting and refracting) and their magnifying powers.

(a) Reflection of light by spherical mirrors. Mirror formula: its derivation; $R = 2f$ for spherical mirrors. Magnification.

(b) Refraction of light at a plane interface, Snell's law; total internal reflection and critical angle; total reflecting prisms and optical fibers. Total reflecting prisms: application to triangular prisms with angle of the prism 30° , 45° , 60° and 90° respectively; ray diagrams for Refraction through a combination of media, ${}_1n_2 \times {}_2n_3 \times {}_3n_1 = 1$, real depth and apparent depth. Simple applications.

(c) Refraction through a prism, minimum deviation and derivation of relation between n , A and δ_{min} . Include explanation of i - δ graph, $i_1 = i_2 = i$ (say) for δ_m ; from symmetry $r_1 = r_2$; refracted ray inside the prism is parallel to the base of the equilateral prism. Thin prism.

Dispersion; Angular dispersion; dispersive power, rainbow - ray diagram (no derivation). Simple explanation.

(d) *Refraction at a single spherical surface; detailed discussion of one case only - convex towards rarer medium, for spherical surface and real image. Derive the relation between n_1 , n_2 , u , v and R . Refraction through thin lenses: derive lens maker's formula and lens formula; derivation of combined focal length of two thin lenses in contact. Combination of lenses and mirrors (silvering of lens excluded) and magnification for lens, derivation for biconvex lens only; extend the results to biconcave lens, plano convex lens and lens immersed in a liquid; power of a lens $P=1/f$ with SI unit dioptre. For lenses in contact $1/F=1/f_1+1/f_2$ and $P=P_1+P_2$. Lens formula, formation of image with combination of thin lenses and mirrors.*

[Any one sign convention may be used in solving numericals].

(e) *Ray diagram and derivation of magnifying power of a simple microscope with image at D (least distance of distinct vision) and infinity; Ray diagram and derivation of magnifying power of a compound microscope with image at D . Only expression for magnifying power of compound microscope for final image at infinity.*

Ray diagrams of refracting telescope with image at infinity as well as at D ; simple explanation; derivation of magnifying power; Ray diagram of reflecting telescope with image at infinity. Advantages, disadvantages and uses. Resolving power of compound microscope.

(ii) Wave Optics

Wave front and Huygen's principle. Proof of laws of reflection and refraction using Huygen's principle. Interference, Young's double slit experiment and expression for fringe width(β), coherent sources and sustained interference of light, Fraunhofer

diffraction due to a single slit, width of central maximum.

(a) *Huygen's principle: wavefronts - different types/shapes of wavefronts; proof of laws of reflection and refraction using Huygen's theory. [Refraction through a prism and lens on the basis of Huygen's theory not required].*

(b) *Interference of light, interference of monochromatic light by double slit. Phase of wave motion; superposition of identical waves at a point, path difference and phase difference; coherent and incoherent sources; interference: constructive and destructive, conditions for sustained interference of light waves [mathematical deduction of interference from the equations of two progressive waves with a phase difference is not required]. Young's double slit experiment: set up, diagram, geometrical deduction of path difference $\Delta x = ds \sin \theta$, between waves from the two slits; using $\Delta x = n\lambda$ for bright fringe and $\Delta x = (n + \frac{1}{2})\lambda$ for dark fringe and $\sin \theta = \tan \theta = y_n/D$ as y and θ are small, obtain $y_n = (D/d)n\lambda$ and fringe width $\beta = (D/d)\lambda$. Graph of distribution of intensity with angular distance.*

(c) *Single slit Fraunhofer diffraction (elementary explanation, qualitative treatment only). Diffraction at a single slit: experimental setup, diagram, diffraction pattern, obtain expression for position of minima, $a \sin \theta_n = n\lambda$, where $n = 1, 2, 3, \dots$ and conditions for secondary maxima, $a \sin \theta_n = (n + \frac{1}{2})\lambda$; distribution of intensity with angular distance; angular width of central bright fringe.*

7. Dual Nature of Radiation and Matter

Wave particle duality; photoelectric effect, Hertz and Lenard's observations; Einstein's photoelectric equation - particle nature of light. Matter waves - wave nature of particles, de-Broglie relation; conclusion from Davisson-Germer experiment (Qualitative only).

(a) *Photo electric effect, quantization of radiation; Einstein's equation*

$E_{max} = h\nu - W_0$; threshold frequency; work function; experimental facts of Hertz and Lenard and their conclusions; Einstein used Planck's ideas and extended it to apply for radiation (light); photoelectric effect can be explained only assuming quantum (particle) nature of radiation. Determination of Planck's constant (from the graph of stopping potential V_s versus frequency f of the incident light). Momentum of photon $p=E/c=h\nu/c=h/\lambda$.

(b) De Broglie hypothesis, phenomenon of electron diffraction (qualitative only). Wave nature of radiation is exhibited in interference, diffraction and polarisation; particle nature is exhibited in photoelectric effect. Dual nature of matter: particle nature common in that it possesses momentum p and kinetic energy KE. The wave nature of matter was proposed by Louis de Broglie, $\lambda=h/p=h/mv$. Davisson and Germer experiment; qualitative description of the experiment and conclusion.

8. Atoms and Nuclei

(i) Atoms

Alpha-particle scattering experiment; Rutherford's atomic model; Bohr's atomic model, energy levels, hydrogen spectrum.

Rutherford's nuclear model of atom (mathematical theory of scattering excluded), based on Geiger - Marsden experiment on α -scattering; nuclear radius r in terms of closest approach of α particle to the nucleus, obtained by equating $\Delta K = \frac{1}{2} mv^2$ of the α particle to the change in electrostatic potential energy ΔU of the system $[U = \frac{2e \times Ze}{4\pi\epsilon_0 r_0}]$ $r_0 \sim 10^{-15} m = 1 \text{ fermi}$; atomic

structure; only general qualitative ideas, including atomic number Z , Neutron number N and mass number A . A brief account of historical background leading to Bohr's theory of hydrogen spectrum; formulae for wavelength in Lyman, Balmer, Paschen, Brackett and Pfund series. Rydberg constant. Bohr's model of H atom, postulates ($Z=1$); expressions for orbital velocity, kinetic

energy, potential energy, radius of orbit and total energy of electron. Energy level diagram, calculation of ΔE , frequency and wavelength of different lines of emission spectra; agreement with experimentally observed values. [Use nm and not \AA for unit of λ].

(ii) Nuclei

Composition and size of nucleus. Mass-energy relation, mass defect; binding energy per nucleon and its variation with mass number; Nuclear reactions, nuclear fission and nuclear fusion.

(a) Atomic masses and nuclear density; Isotopes, Isobars and Isotones – definitions with examples of each. Unified atomic mass unit, symbol u , $1u = 1/12$ of the mass of ^{12}C atom = $1.66 \times 10^{-27} \text{ kg}$. Composition of nucleus; mass defect and binding energy, $BE = (\Delta m) c^2$. Graph of $BE/\text{nucleon}$ versus mass number A , special features - less $BE/\text{nucleon}$ for light as well as heavy elements. Middle order more stable [see fission and fusion] Einstein's equation $E=mc^2$. Calculations related to this equation; mass defect/binding energy, mutual annihilation and pair production as examples.

(b) Nuclear Energy

Theoretical (qualitative) prediction of exothermic (with release of energy) nuclear reaction, in fusing together two light nuclei to form a heavier nucleus and in splitting heavy nucleus to form middle order (lower mass number) nuclei, is evident from the shape of BE per nucleon versus mass number graph. Also calculate the disintegration energy Q for a heavy nucleus ($A=240$) with $BE/A \sim 7.6 \text{ MeV}$ per nucleon split into two equal halves with $A=120$ each and $BE/A \sim 8.5 \text{ MeV}/\text{nucleon}$; $Q \sim 200 \text{ MeV}$. Nuclear fission: Any one equation of fission reaction. Chain reaction-controlled and uncontrolled; nuclear reactor and nuclear bomb. Main parts of a nuclear reactor including their functions - fuel elements, moderator,

control rods, coolant, casing; criticality; utilization of energy output - all qualitative only. Fusion, simple example of $4^1\text{H} \rightarrow ^4\text{He}$ and its nuclear reaction equation; requires very high temperature $\sim 10^6$ degrees; difficult to achieve; hydrogen bomb; thermonuclear energy production in the sun and stars. [Details of chain reaction not required].

9. Electronic Devices

(i) Semiconductor Electronics: Materials, Devices and Simple Circuits. Energy bands in conductors, semiconductors and insulators (qualitative ideas only). Intrinsic and extrinsic semiconductors. P and n type, p-n junction.

(ii) Semiconductor diode: I-V characteristics in forward and reverse bias, diode as a rectifier; Special types of junction diodes: LED, photodiode and solar cell and Zener diode and its characteristics, Zener diode as a voltage regulator.

(a) *Energy bands in solids; energy band diagrams for distinction between conductors, insulators and semi-conductors - intrinsic and extrinsic; electrons and holes in semiconductors.*

Elementary ideas about electrical conduction in metals [crystal structure not included]. Energy levels (as for hydrogen atom), 1s, 2s, 2p, 3s, etc. of an isolated atom such as that of copper; these split, eventually forming 'bands' of energy levels, as we consider solid copper made up of a large number of isolated atoms, brought together to form a lattice; definition of energy bands - groups of closely spaced energy levels separated by band gaps called forbidden bands. An idealized representation of the energy bands for a conductor, insulator and semiconductor; characteristics, differences; distinction between conductors, insulators and semiconductors on the basis of energy bands, with examples; qualitative discussion only; energy gaps (eV) in typical substances (carbon, Ge, Si); some electrical properties of semiconductors. Majority and minority charge carriers - electrons and holes;

intrinsic and extrinsic, doping, p-type, n-type; donor and acceptor impurities.

(b) *Junction diode and its symbol; depletion region and potential barrier; forward and reverse biasing, V-I characteristics and numericals; half wave and a full wave rectifier. Simple circuit diagrams and graphs, function of each component in the electric circuits, qualitative only. [Bridge rectifier of 4 diodes not included]; elementary ideas on solar cell, photodiode and light emitting diode (LED) as semi conducting diodes. Importance of LED's as they save energy without causing atmospheric pollution and global warming. Zener diode, V-I characteristics, circuit diagram and working of Zener diode as a voltage regulator.*

PAPER II

PRACTICAL WORK- 15 Marks

The experiments for laboratory work and practical examinations are mostly from two groups:

- (i) experiments based on ray optics and
- (ii) experiments based on current electricity.

The main skill required in group (i) is to remove parallax between a needle and the real image of another needle.

In group (ii), understanding circuit diagram and making connections strictly following the given diagram is very important. Polarity of cells and meters, their range, zero error, least count, etc. should be taken care of.

A graph is a convenient and effective way of representing results of measurement. It is an important part of the experiment.

There will be one graph in the Practical question paper.

Candidates are advised to read the question paper carefully and do the work according to the instructions given in the question paper. Generally they are not expected to write the procedure of the experiment, formulae, precautions, or draw the figures, circuit diagrams, etc.

Observations should be recorded in a tabular form.

Record of observations

- All observations recorded should be consistent with the least count of the instrument used (e.g.

focal length of the lens is 10.0 cm or 15.1cm but **10 cm is a wrong record.**)

- All observations should be recorded with correct units.

Graph work

Students should learn to draw graphs correctly noting all important steps such as:

- (i) Title
- (ii) Selection of origin (should be marked by two coordinates, example 0,0 or 5,0, or 0,10 or 30,5; **Kink is not accepted**).
- (i) The axes should be labelled according to the question
- (ii) Uniform and convenient scale should be taken and the units given along each axis (one small division = 0.33, 0.67, 0.66, etc. should not to be taken)
- (iii) Maximum area of graph paper (**at least 60% of the graph paper along both the axes**) should be used.
- (iv) Points should be plotted with great care, marking the points plotted with (should be a circle with a dot) \square or \otimes . A blob \bullet) is a misplot.
- (v) The best fit straight line should be drawn. The best fit line does not necessarily have to pass through all the plotted points and the origin. While drawing the best fit line, **all experimental points must be kept on the line or symmetrically placed on the left and right side of the line**. The line should be continuous, thin, uniform and extended beyond the extreme plots.
- (vi) The intercepts must be read carefully. Y intercept i.e. y_0 is that value of y when $x = 0$. Similarly, X intercept i.e. x_0 is that value of x when $y=0$. **When x_0 and y_0 are to be read, origin should be at (0, 0).**

Deductions

- (i) The slope 'S' of the best fit line must be found taking two distant points (**using more than 50% of the line drawn**), which are not the plotted points, using $S = \frac{y_2 - y_1}{x_2 - x_1} = \frac{\Delta y}{\Delta x}$. Slope S must be calculated upto proper decimal place or

significant figures as specified in the question paper.

- (ii) All calculations should be rounded off upto proper decimal place or significant figures, as specified in the question papers.

NOTE:

Short answer type questions may be set from each experiment to test understanding of theory and logic of steps involved.

Given below is a list of required experiments. Teachers may add to this list, keeping in mind the general pattern of questions asked in the annual examinations.

Students are required to have completed all experiments from the given list (excluding demonstration experiments):

1. To find focal length of a convex lens by using u-v method (no parallax method)

Using a convex lens, optical bench/metre scales and two pins, obtain the positions of the images for various positions of the object; $f < u < 2f$, $u \sim 2f$, and $u > 2f$.

Draw the following set of graphs using data from the experiments -

- (i) v against u . It will be a curve.

(ii) Magnification $\left(m = \frac{v}{u} \right)$ against v which is a straight line and to find focal length by intercept.

(iii) $y = (100/v)$ against $x = (100/u)$ which is a straight line and find f by intercepts.

2. To find f of a convex lens by displacement method.

3. To determine the focal length of a given convex lens with the help of an auxiliary convex lens.

4. To determine the focal length of a concave lens, using an auxiliary convex lens, not in contact and plotting appropriate graph.

5. To determine focal length of concave mirror by using two pins (by u-v method).

6. To determine the refractive index of a liquid by using a convex lens and a plane mirror.

7. To determine the focal length of a convex mirror using convex lens.
8. Using a metre bridge, determine the resistance of about 100 cm of (constantan) wire. Measure its length and radius and hence, calculate the specific resistance of the material. Verify Ohm's law for the given unknown resistance (a 60 cm constantan wire), plotting a graph of potential difference versus current. Also calculate the resistance per cm of the wire from the slope of the graph and the length of the wire.
9. To determine the internal resistance of a cell by a potentiometer.
10. From a potentiometer set up, measure the fall in potential (i.e. pd) for increasing lengths of a constantan wire, through which a steady current is flowing; plot a graph of pd (V) versus length (l). Calculate the potential gradient of the wire and specific resistance of its material. Q (i) Why is the current kept constant in this experiment? Q (ii) How can you increase the sensitivity of the potentiometer? Q (iii) How can you use the above results and measure the emf of a cell?
11. To verify the laws of combination of resistances (series and parallel) using metre bridge.

Demonstration Experiments (The following experiments are to be demonstrated by the teacher):

1. To convert a given galvanometer into (a) an ammeter of range, say 2A and (b) a voltmeter of range 4V.
2. To study I-V characteristics of a semi-conductor diode in forward and reverse bias.
3. To determine refractive index of a glass slab using a traveling microscope.
4. Identification of diode, LED, transistor, IC, resistor, capacitor from mixed collection of such items.
5. Use of multimeter to (i) identify base of transistor, (ii) distinguish between npn and pnp type transistors, (iii) see the unidirectional flow of current in case of diode and an LED, (iv) check whether a given electronic component (e.g. diode, transistors, IC) is in working order.
6. Charging and discharging of a capacitor.

PROJECT WORK AND PRACTICAL FILE – 15 marks

Project Work – 10 marks

The Project work is to be assessed by a Visiting Examiner appointed locally and approved by the Council.

All candidates will be required to do **one** project involving some physics related topic/s under the guidance and regular supervision of the Physics teacher.

Candidates should undertake any **one** of the following types of projects:

- Theoretical project
- Working Model
- Investigatory project (by performing an experiment under supervision of a teacher)

Candidates are to prepare a technical report including title, abstract, some theoretical discussion, experimental setup, observations with tables of data collected, graph/chart (if any), analysis and discussion of results, deductions, conclusion, etc. The teacher should approve the draft, before it is finalised. The report should be kept simple, but neat and elegant. Teachers may assign or students may choose **any one** project of their choice.

Suggested Evaluation Criteria for Theory Based Projects:

▪ Title of the Project
▪ Introduction
▪ Contents
▪ Analysis/ material aid (graph, data, structure, pie charts, histograms, diagrams, etc.)
▪ Originality of work (the work should be the candidates' original work,)
▪ Conclusion/comments

Suggested Evaluation Criteria for Model Based Projects:

▪ Title of the Project
▪ Model construction
▪ Concise Project report

Suggested Evaluation Criteria for Investigative Projects:

▪ Title of the Project
▪ Theory/principle involved
▪ Experimental setup
▪ Observations calculations/deduction and graph work
▪ Result/ Conclusions

Practical File – 5 marks

The Visiting Examiner is required to assess the candidates on the basis of the Physics practical file maintained by them during the academic year.

CHEMISTRY (862)

CLASS XII

There will be two papers in the subject:

Paper I: Theory - *3 hours ... 70 marks*

Paper II: Practical: *3 hours ... 15 marks*

Project Work *... 10 marks*

Practical File *... 5 marks*

PAPER I - THEORY: 70 Marks

S.No.	UNIT	TOTAL WEIGHTAGE
1.	Solutions	Physical Chemistry 25 Marks
2.	Electrochemistry	
3.	Chemical Kinetics	
4.	d -and f -Block Elements	Inorganic Chemistry 14 Marks
5.	Coordination Compounds	
6.	Haloalkanes and Haloarenes	Organic Chemistry 31 Marks
7.	Alcohols, Phenols and Ethers	
8.	Aldehydes, Ketones and Carboxylic Acids	
9.	Organic Compounds containing Nitrogen	
10.	Biomolecules	
Total		70 Marks

PAPER I-THEORY – 70 Marks

1. Solutions

Study of concentration of solutions of solids in liquids, liquid in liquid, solubility of gases in liquids, solid solutions, Colligative properties - Raoult's law of relative lowering of vapour pressure (1st & 2nd), elevation of boiling point, depression of freezing point, osmotic pressure. Use of colligative properties in determining molecular masses of solutes, abnormal molecular mass association and dissociation, van't Hoff factor.

Normality, molality, molarity, mole fraction, ppm, as measures of concentration. Definition of the above with examples. Simple problems based on the above.

- (i) *Solubility of gases in liquids – Henry's Law, simple numericals based on the above.*
- (ii) *Raoult's Law for volatile solutes and non-volatile solutes, ideal solution, non-ideal solution. Azeotropic mixtures – definition, types, graphical representation, fractional distillation with examples.*
- (iii) *Colligative properties – definition and examples, and its use in determination of molecular mass.*
 - (a) *Relative lowering of vapour pressure: Definition and mathematical expression of Raoult's Law. Determination of relative molecular mass by measurement of lowering of vapour pressure.*
 - (b) *Depression in freezing point: molal depression constant (cryoscopic constant) – definition and mathematical expression (derivation included).*
 - (c) *Elevation in boiling point method: molal elevation constant (ebullioscopic constant) definition and mathematical expression (derivation included).*
 - (d) *Osmotic pressure: definition and explanation. Natural and chemical semipermeable membranes, reverse osmosis, isotonic, hypotonic and hypertonic solutions. Comparison between diffusion and osmosis. Application of osmotic pressure in the determination of relative molecular mass.*

van't Hoff- Boyle's Law, van't Hoff – Charles' Law, van't Hoff - Avogadro's law.

- (e) *Abnormal molecular mass: Dissociation and Association with suitable examples*
- (f) *van't Hoff factor for the electrolytes which dissociate and the molecules which associate in solution. Modification of the formula of colligative properties based on van't Hoff factor. Simple problems. Calculation of degree of dissociation and association. Experimental details not required.*

Numerical problems based on all the above methods. Experimental details not required.

2. Electrochemistry

Electrolytic and electrochemical cells. Redox reactions in electrochemical cells. Electromotive Force (emf) of a cell, standard electrode potential, Nernst equation and its application to chemical cells. Relation between Gibbs energy change and emf of a cell.

Conductance in electrolytic solutions, specific, equivalent and molar conductivity, variations of conductivity with concentration, graphs; Kohlrausch's Law of electrolysis and Faraday's Laws of electrolysis. Dry cell and lead accumulator, fuel cells, corrosion.

- (i) *Electrochemical cells: introduction, redox reactions (principle of oxidation and reduction in a cell).*
- (ii) *Galvanic cells - introduction; representation, principle – oxidation reduction. Mechanism of production of electric current in a galvanic cell.*
- (iii) *Measurement of potential. Single electrode potentials.*
Standard hydrogen electrode (E°) - definition, preparation, application and limitations.

Standard electrode potential - Measurement of standard electrode potential of Zn^{++}/Zn , Cu^{++}/Cu , half cell (using standard hydrogen electrode).

Cell notation – representation.

Factors affecting electrode potential with explanation - main emphasis on the

temperature, concentration and nature of the electrode.

(iv) *Electrochemical series. Its explanation on the basis of standard reduction potential. Prediction of the feasibility of a reaction.*

(v) *Nernst equation and correlation with the free energy of the reaction with suitable examples. Prediction of spontaneity of a reaction based on the cell emf. Numericals on standard electrode potential of half-cells, cell emf, relationship between free energy and equilibrium constant, standard electrode potential and free energy.*

(vi) *Comparison of metallic conductance and electrolytic conductance. Relationship between conductance and resistance. Specific resistance and specific conductance. Cell constant: Calculation of cell constant. Meaning of equivalent conductance. Meaning of molar conductance. General relationship between specific conductance, molar conductance and equivalent conductance (units and graphs). Units, numericals. Molar conductance of a weak electrolyte at a given concentration and at infinite dilution. Kohlrausch's Law – definition, applications and numericals.*

(vii) *Faraday's laws of Electrolysis. Faraday's First Law of electrolysis. Statement, mathematical form. Simple problems. Faraday's Second Law of electrolysis: Statement, mathematical form. Simple problems. Relation between Faraday, Avogadro's number and charge on an electron. $F = N_A e$ should be given (no details of Millikan's experiment are required).*

(viii) *Batteries: Primary and Secondary Cells: Leclanche cell, mercury cell, Lead storage battery and fuel cell – structure, reactions and uses.*

(ix) *Corrosion: Concept, mechanism of electrochemical reaction, factors affecting it and its prevention.*

3. Chemical Kinetics

Meaning of Chemical Kinetics – slow and fast reactions. Rate of a reaction - average and instantaneous rate (graphical representation). Factors affecting rate of reaction: surface area, nature of reactants, concentration, temperature, catalyst and radiation. Order and molecularity of a reaction, rate law and specific rate constant. Integrated rate equations and half-life (only for zero and first order reactions), concept of collision theory (elementary idea, no mathematical treatment). Concept of threshold and activation energy, Arrhenius equation.

(i) *Meaning of chemical kinetics, Scope and importance of Kinetics of the reaction, slow and fast reactions – explanation in terms of bonds.*

(ii) *Rate of Reaction: definition, representation of rate of reaction in terms of reactants and products, determination of rate of reactions graphically, instantaneous and average rate of reaction. Factors affecting rate of reaction.*

(iii) *Law of mass Action: statement and meaning of active mass. Explanation with an example – general reactions.*

(iv) *Effect of concentration of reactants on the rate of a reaction: Qualitative treatment, based on the law of mass Action, statement of rate law, General rate equation – $\text{Rate} = k(\text{concentration of the reactant})^n$, where k is rate constant and n is the order of the reaction, relationship between the rate of the reaction with rate constant with respect to various reactants.*

(v) *Order of a reaction: meaning, relation between order and stoichiometric coefficients in balanced equations, order as an experimental quantity, rate equation for zero order reaction and its unit, mathematical derivation of rate equation for first order reaction, characteristics of first order reaction – rate constant is independent of the initial concentration, units to be derived,*

definition of half-life period, derivation of expression of half-life period from first order rate equation.

Problems based on first order rate equation and half-life period.

(vi) *Molecularity of the reaction: Meaning – physical picture, Relation between order, molecularity and the rate of a reaction, Differences between order and molecularity of a reaction.*

(vii) *The concept of energy: Exothermic and endothermic reactions, concept of energy barrier, threshold and activation energy, formation of activated complex, effect of catalyst on activation energy and reaction rate.*

(viii) *Collision Theory: Condition for a chemical change – close contact, particles should collide. Collisions to be effective – optimum energy and proper orientation during collision. Energy barrier built-up when the collision is about to take place, Activated complex formation, difference in energy of the reactant and the product – exothermic and endothermic reactions with proper graphs and labelling.*

(ix) *Mechanism of the reaction: meaning of elementary reaction, meaning of complex and overall reaction, explanation of the mechanism of the reaction, slowest step of the reaction. Relationship between the rate expression, order of reactants and products at the rate-determining step, units of rate constant – explanation with suitable examples.*

(x) *Effect of temperature on the rate constant of a reaction: Arrhenius equation – $K=Ae^{-E_a/RT}$, Meaning of the symbols of Arrhenius equation, related graph, evaluation of E_a and A from the graph, meaning of slope of the graph, conversion from exponential to log form of the equation, relationship between the increase in temperature and the number of collisions. Numerical based on Arrhenius equation.*

4. d- and f- Block Elements

Position in the periodic table, occurrence, electronic configuration and characteristics of transition metals, general trends in properties of the 3d-series of transition metals - metallic character, ionisation enthalpy, oxidation states, ionic radii, colour of ions, catalytic property, magnetic properties, interstitial compounds, alloy formation, preparation and properties of $K_2Cr_2O_7$ and $KMnO_4$.

Lanthanoids and actinoids.

(i) *d-Block: 3d, 4d and 5d series*

Study in terms of metallic character, atomic and ionic radii, ionisation enthalpy, oxidisation states, variable valency, formation of coloured compounds, formation of complexes, alloy formation.

(ii) *f-Block: 4f and 5f series*

Electronic configuration, atomic and ionic radii, oxidisation states, formation of coloured compounds, formation of complexes, alloy formation. Lanthanoid contraction and its consequences. Chemical reactivity – with oxygen, hydrogen, halogen, sulphur, nitrogen, carbon and water.

Actinoids - oxidation states and comparison with lanthanoids.

(iii) *Potassium permanganate: structure, shape, equation of extraction from pyrolusite ore, its oxidising nature in acidic, basic and neutral medium, use in redox titration.*

Oxidising nature in acidic $[FeSO_4, (COOH)_2 \cdot 2H_2O, KI]$, basic (KI) and neutral (H_2S) mediums to be done.

(iv) *Potassium dichromate: structure, shape, equation of extraction from chromite ore and its use in titration. Oxidising nature in acidic, basic and neutral medium, use in redox titration. Interconversion of chromate and dichromate ion (effect of pH).*

5. Coordination Compounds

Concept of complexes, definition of ligands, coordination number, oxidation number. IUPAC nomenclature of mononuclear coordination

compounds. Isomerism (structural and stereo). Bonding, Werner's theory, VBT and CFT. Colour, magnetic properties and shapes. Importance of coordination compounds (in qualitative analysis, extraction of metals and biologicals system).

- (i) *Definition of coordination compounds / complex compounds, differences with a double salt, study of ligands – mono-, bi-, tri-, tetra-, penta-, hexa- and polydentate, chelating ligands, definition of coordination number, its calculation for a complex coordination sphere, study of oxidation state of an element in a complex, its calculation, IUPAC rules of nomenclature of coordination compounds.*
- (ii) *Isomerism – structural, stereo types and examples.*
- (iii) *Valence bond theory of coordination compounds – examples of formation of inner orbital and outer orbital complexes (high and low spin, octahedral, tetrahedral and square planar), prediction of magnetic character.*
- (iv) *Crystal field theory – crystal field splitting in tetra and octahedral systems. Explanation of colour and magnetic character.*
- (v) *Stability of coordination compounds (explain stability on the basis of magnitude of K) as mentioned above).*
- (vi) *Importance and uses.*

6. Haloalkanes and Haloarenes.

Haloalkanes: General formula, nomenclature and classification. Nature of C–X bond, physical and chemical properties, mechanism of substitution reactions, optical rotation.

Haloarenes: Basic idea, nature of C–X bond, substitution reactions (directive influence of halogen in monosubstituted compounds only).

Uses and environmental effects of - dichloromethane, trichloromethane, tetrachloromethane, iodoform, freons and DDT.

Nature of C-X bond

Naming the halogen derivatives of alkanes by using common system and IUPAC system for mono, di and tri-halo derivatives.

Preparation of haloalkanes from:

- *Alkane and halogen.*
- *Alkene and hydrogen halide.*
- *Alcohols with PX_3 , PCl_5 and $SOCl_2$.*
- *Halide exchange method (Finkelstein and Swarts)*
- *Silver salt of fatty acids (Hunsdiecker).*

Physical properties: State, melting point, boiling point and solubility.

Chemical properties: nucleophilic substitution reactions (S_N1 , S_N2 mechanism in terms of primary, secondary and tertiary halides) Reaction with: sodium hydroxide, water, sodium iodide, ammonia, primary amine, secondary amine, potassium cyanide, silver cyanide, potassium nitrite, silver nitrite, silver salt of fatty acid and lithium-aluminium hydride.

Elimination reaction (Saytzeff's rule) / β elimination.

Reaction with metals: sodium and magnesium (Wurtz's reaction, Grignard's reagent preparation).

Chloroform and iodoform: preparation and properties.

Structure of freons.

Preparation of haloarenes by Sandmeyer's and Gattermann's reaction, by electrophilic substitution.

Physical properties: State, melting point, boiling point and solubility.

Chemical properties:

- *Electrophilic substitution (chlorination nitration and sulphonation) with mechanism.*
- *Nucleophilic substitution (replacement of chlorine with $-OH$, $-NH_2$) with mechanism.*
- *Reduction to benzene.*
- *Wurtz-Fittig reaction.*
- *Fittig reaction.*

- Addition reaction with magnesium (formation of Grignard reagent).
- Structure of DDT.

7. Alcohols, Phenols and Ethers

Alcohols: Classification, general formula, structure and nomenclature. Methods of preparation, physical and chemical properties (of primary alcohols only), identification of primary, secondary and tertiary alcohols, mechanism of dehydration, uses with special reference to methanol and ethanol.

(i) Classification into monohydric, dihydric and polyhydric alcohols, general formulae, structure and nomenclature of alcohols. Difference between primary, secondary and tertiary alcohols in terms of structure, physical properties and chemical properties.

(ii) Methods of preparation:

- Hydration of Alkenes – direct hydration, indirect hydration, hydroboration oxidation.
- From Grignard's reagent.
- Hydrolysis of alkyl halides.
- Reduction of carbonyl compounds.
- From primary amines.

Manufacture of methanol by Bosch process and ethanol by fermentation of carbohydrates, chemical equations required (only outline of the method of manufacture, detail not required).

Properties:

- Acidic nature of alcohols.
- Reaction with sodium.
- Esterification with mechanism.
- Reaction with hydrogen halides.
- Reaction with PCl_3 , PCl_5 , and $SOCl_2$.
- Reaction with acid chlorides and acid anhydrides
- Oxidation.
- Dehydration with mechanism.

Uses of alcohols.

(iii) Conversion of one alcohol into another.

(iv) Distinction between primary, secondary and tertiary alcohols by Lucas' Test.

Phenols: Classification and nomenclature. Methods of preparation, physical and chemical properties, acidic nature of phenol, electrophilic substitution reactions, uses of phenols.

Preparation of phenol from diazonium salt, chlorobenzene (Dow's process) and from benzene sulphonic acid.

Manufacture from Cumene.

Physical properties: state and solubility.

Chemical properties:

- Acidic character of phenol.
- Reaction with sodium hydroxide.
- Reaction with sodium.
- Reaction with zinc.
- Reaction with acetyl chloride and acetic anhydride.
- Reaction with phosphorus penta chloride.
- Bromination, nitration and sulphonation (Electrophilic substitution reactions).
- Kolbe's reaction (formation of salicylic acid).
- Reimer – Tiemann reaction
- Test for phenol – $FeCl_3$ test, azo dye test.

Aliphatic Ethers: General formula, structure and nomenclature. Methods of preparation, physical and chemical properties, uses.

Ethers: structure of ethereal group.

Preparation from alcohol (Williamson's synthesis).

Physical properties: state, miscibility.

Chemical properties:

- Reaction with chlorine.
- Oxidation (peroxide formation).
- Reaction with HI .
- Reaction with PCl_5 .

Aryl ethers

Physical properties – state and solubility.

Chemical properties – preparation of anisole (Williamson's synthesis), electrophilic substitution (halogenation, nitration and Friedel-Crafts reaction.)

Uses of ether.

8. Aldehydes, Ketones and Carboxylic Acids

Aldehydes and Ketones: Nomenclature, structure of methods of preparation of aldehydes and ketones, physical and chemical properties, mechanism of nucleophilic addition, reactivity of alpha hydrogen in aldehydes and uses.

Preparation:

- *From alcohol.*
- *From alkenes (ozonolysis).*
- *From alkynes (hydration).*
- *From acid chlorides (Rosenmund's reduction, reaction with dialkyl cadmium).*
- *From calcium salt of carboxylic acids.*
- *From nitriles (Stephen reaction, Grignard's reagent).*
- *From esters.*

Physical properties – state and boiling point.

Chemical properties:

- *Nucleophilic addition reactions with mechanism (ammonia and its derivatives, HCN, NaHSO₃ and Grignard's reagent).*
- *Oxidation reactions, iodoform reaction.*
- *Reduction: reduction to alcohol and alkanes (Clemmensen's reduction, Wolff-Kishner reduction, Red phosphorus and HI).*
- *Base catalysed reactions (with mechanism): Aldol condensation, cross Aldol condensation, Cannizzaro's reaction.*

Tests: difference between formaldehyde and acetaldehyde; aldehydes and ketones.

Uses of aldehydes and ketones.

Aromatic aldehyde (Benzaldehyde)

Lab preparation from toluene by oxidation with chromyl chloride.

Physical properties: state and stability.

Chemical properties:

- *Oxidation and reduction.*
- *Nucleophilic addition reaction (hydrogen cyanide and sodium bisulphite).*
- *Reactions with ammonia and its derivatives (hydroxyl amine, hydrazine and phenyl hydrazine).*
- *Reaction with phosphorus pentachloride.*
- *Cannizzaro reaction.*
- *Benzoin condensation.*
- *Perkin's reaction.*
- *Electrophilic substitution - halogenation, nitration and sulphonation.*

Test: distinction between aromatic and aliphatic aldehydes.

Uses of benzaldehyde.

Carboxylic Acids: Classification, general formula and structure of carboxylic group. Nomenclature, acidic nature, methods of preparation, physical and chemical properties and uses.

Classification of mono and di carboxylic acids with examples.

Preparation of aliphatic and aromatic carboxylic acid:

- *From alcohols, aldehydes.*
- *From nitriles.*
- *From Grignard's reagent.*

Physical properties: state, boiling point and solubility.

Chemical properties:

- *Acidic character: (aliphatic, aromatic carboxylic acids with the effect of substituents on the acidic character – to be dealt with in detail)*

- Reaction with active metals, alkalies, carbonates and bicarbonates,
- Formation of acid derivatives.
- Decarboxylation (chemical and Kolbe's electrolytic reaction).
- HVZ reactions.
- Substitution of benzene ring (meta directive effect of carboxylic acid group) nitration and sulphonation.

Tests for acids: formic acid, acetic acid and benzoic acid.

Uses of formic acid, acetic acid and benzoic acid.

9. Organic compounds containing Nitrogen

Aliphatic Amines: General formula and, classification of amines. Structure of the amino group, nomenclature. Methods of preparation, physical and chemical properties, uses, identification of primary, secondary and tertiary amines.

- **Amines**

Nomenclature, classification with examples, structure, general formula.

Methods of preparation:

- From alcohol.
- From alkyl halide.
- From cyanide.
- From amide (Hofmann's degradation).
- From nitro compounds.
- Gabriel phthalimide Synthesis.

Physical properties: comparison between primary, secondary and tertiary amines in terms of – state, solubility, boiling point (hydrogen bonding), comparison with alcohols.

Chemical properties:

- Basic character of amines – comparison between primary, secondary and tertiary alkyl amines/ ammonia/ aniline. Effect of substituents on the basic strength of aniline

- Alkylation and acylation with mechanism.
- Reaction with nitrous acid.
- Carbylamine reaction.

Distinction between primary, secondary and tertiary amines (Hinsberg's Test).

Aniline

Preparation reduction of nitrobenzene.

Physical properties – state, solubility and boiling point.

Chemical properties:

- Reaction with HCl and H_2SO_4 .
- Acetylation, alkylation.
- Benzoylation.
- Carbylamine reaction.
- Diazotisation.
- Electrophilic substitution (bromination, nitration and sulphonation).

Tests for aniline.

Uses of aniline.

Cyanides and Isocyanides

Methods of preparation:

Cyanides:

- From alkyl halide.
- From amide.

Isocyanides:

- From alkyl halide.

From primary amines

Diazonium salts: Preparation, chemical reactions and importance in synthetic organic chemistry.

Preparation from aniline;

Properties: Sandmeyer's reaction, Gattermann reaction and Balz – Scheimann reaction, replacement of diazo group by – H , $-OH$, $-NO_2$, coupling reaction with phenol and aniline.

10. Biomolecules

Carbohydrates – Definition, Classification (aldoses and ketoses), monosaccharides (glucose and fructose), D-L configuration

oligosaccharides (sucrose, lactose, maltose), polysaccharides (starch, cellulose, glycogen); Importance of carbohydrates.

Carbohydrates: definition, classification - mono (aldose, ketose), oligo (di, tri, tetra saccharides) and polysaccharides with examples: reducing sugars and non-reducing sugars – examples and uses.

Establishment of structures for glucose and fructose (open and cyclic) heating with HI, reaction with hydroxylamine, bromine water, acetic anhydride, nitric acid and phenyl hydrazine.

Test for glucose and fructose (bromine water test with equation).

Disaccharides – structures of sucrose, maltose and lactose (glycosidic linkage).

Polysaccharides – starch, cellulose, glycogen.

Proteins – structural units of proteins. Basic idea of - amino acids, peptide bond, polypeptides, proteins, structure of proteins - primary, secondary, tertiary structure and quaternary structures (qualitative idea only), denaturation of proteins. Enzymes, hormones - elementary idea only.

Proteins: Amino acids – general structure, classification and zwitter ion formation. Isoelectric point.

Classification of proteins on the basis of molecular shape; primary, secondary, tertiary and quaternary, structures of proteins, denaturation of proteins. (Definitions only. Details and diagrams are not required).

Vitamins - Classification and functions.

Vitamins A, B, C, D, E and K: classification (fat soluble and water soluble), deficiency diseases. (Chemical names and structures are not required).

Nucleic Acids - DNA and RNA.

Nucleic acids: basic unit – purine and pyrimidine, DNA – structure (double helical), RNA (No chemical structure required). Differences between DNA and RNA.

PAPER II

PRACTICAL WORK – 15 Marks

Candidates are required to complete the following experiments:

1. Titrations

Oxidation-reduction titrations: potassium manganate (VII) / ammonium iron (II) sulphate; potassium manganate (VII) / oxalic acid.

The candidate may be required to determine the percentage purity of a compound and the number of molecules of water of crystallization in hydrated salts. In such experiments sufficient working details including recognition of the end point will be given.

Candidates will be required to calculate:

- Molarity
- Concentration in grams L^{-1} / molecular mass
- Number of molecules of water of crystallisation/percentage purity.

NOTE: Molarity must be calculated upto 4 decimal places at least, in order to avoid error.

OBSERVATION TABLE

S. No.	(A)	(B)	(B – A)
	Initial burette reading (ml)	Final burette reading (ml)	Difference (ml)
1			
2			
3			

- Concordant reading is to be used for titre value. Concordant reading is two consecutive values which are exactly the same. Average will not be accepted as titre value.
- The table is to be completed in ink only. Pencil is not to be used.
- Overwriting will not be accepted in the tabular column.

Observations:

- Pipette size (should be same for all the candidates at the centre).
- Titre value (concordant value).

2. Study of the rate of reaction

The candidates will be required, having been given full instructions, to carry out an experiment on the rate of reaction, e.g. reaction between sodium thiosulphate and hydrochloric acid (using different concentrations for either), magnesium and dil. sulphuric acid/ dil. hydrochloric acid (using different concentrations).

- Graph of volume vs. time and its interpretation.
- Relationship between concentration and rate, volume and rate and time and rate.

3. Identification of the following compounds and functional groups based on observations

- Alcoholic group - glycerol
- Aldehyde group- formaldehyde
- Ketonic group – acetone
- Carboxylic group – benzoic acid
- Amino group - aniline

***Please Note: Carbylamine and acrolein tests should not be performed.**

The student should learn to differentiate between colours, solution, ring and precipitate.

4. Characteristic tests of carbohydrates and proteins

- Carbohydrates – glucose
- Proteins – powdered milk

Identification should be of 'Carbohydrate' and 'Protein' not of individual substances.

5. Experiments related to pH change using pH paper or universal indicator.

- Determination of pH of some solutions obtained from fruit juice, solutions of known and varied concentrations of acids, bases and salts.
- Comparison of pH of the solutions of strong and weak acids of the same concentration.

Use of universal indicator/pH paper must be taught to the students.

6. Electrochemistry

Setting up a simple voltaic cell.

Variation of cell potential in $Zn/Zn^{2+}/Cu^{2+}/Cu$ with change in concentration of electrolyte ($CuSO_4$, $ZnSO_4$) at room temperature.

7. Qualitative analysis

Qualitative analysis: identification of single salt containing one anion and one cation:

Anions: CO_3^{2-} , NO_2^- , S^{2-} , SO_3^{2-} , SO_4^{2-} , NO_3^- , CH_3COO^- , Cl^- , Br^- , I^- , $C_2O_4^{2-}$, PO_4^{3-} .

Cations: NH_4^+ , Pb^{2+} , Cu^{2+} , Al^{3+} , Fe^{3+} , Zn^{2+} , Mn^{2+} , Ni^{2+} , Co^{2+} , Ba^{2+} , Sr^{2+} , Ca^{2+} , Mg^{2+} .

NOTE:

Chromyl chloride test not to be performed.

For wet test of anions, sodium carbonate extract must be used (except for carbonate).

(Insoluble salts such as lead sulphate, barium sulphate, calcium sulphate, strontium sulphate will not be given).

Anions: Dilute acid group – CO_3^{2-} , NO_2^- , S^{2-} , SO_3^{2-}

Concentrated Acid Group – NO_3^- , Cl^- , Br^- , I^- , CH_3COO^- .

Special Group - SO_4^{2-} , PO_4^{3-} , $C_2O_4^{2-}$.

Cations: Group Zero: NH_4^+

Group I: Pb^{2+}

Group II: Cu^{2+} , Pb^{2+}

Group III: Al^{3+} , Fe^{3+}

Group IV: Zn^{2+} , Mn^{2+} , Ni^{2+} , Co^{2+}

Group V: Ba^{2+} , Sr^{2+} , Ca^{2+}

Group VI: Mg^{2+}

NOTE:

- Formal analytical procedure is required for Qualitative Analysis.
- Specific solvent for O.S. to be used;
- Before adding Group III reagents to the filtrate of Group II, H_2S must be removed followed by boiling with conc. Nitric acid.
- The right order for buffer (NH_4Cl and NH_4OH) must be used.
- The flame test with the precipitate obtained in Group V for Ba^{2+} , Sr^{2+} , Ca^{2+} will also be accepted as a confirmatory test.

For wet test of anions, sodium carbonate extract must be used (except for carbonate).

PATTERN OF CHEMISTRY PRACTICAL PAPER

Questions in the practical paper will be set as follows:

Question 1	Volumetric Analysis
Question 2	Any one or a combination of the following experiments: <ul style="list-style-type: none"> • Study of the rate of reaction. • Identification of the organic compounds and functional groups based on observations. • Characteristic tests of carbohydrates and proteins. • Experiments related to pH determination using pH paper or universal indicator. • Electrochemistry.
Question 3	Qualitative Analysis (single salt).

PROJECT WORK AND PRACTICAL FILE - 15 Marks

Project Work – 10 Marks

The project work is to be assessed by a Visiting Examiner appointed locally and approved by the Council.

The candidate is to creatively execute **one** project/assignment on an aspect of Chemistry. Teachers may assign or students may select a topic of their choice. Following is only a suggestive list of projects.

Suggested Evaluation criteria for Project Work:

- Introduction / purpose
- Contents
- Analysis/ material aid (graph, data, structure, pie charts, histograms, diagrams, etc.)
- Presentation
- Bibliography

Suggested Assignments:

NOTE: According to the recommendation of International Union of Pure and Applied Chemistry (IUPAC), the groups are numbered from 1 to 18 replacing the older notation of groups IA VIIA, VIII, IB VIIIB and 0. However, for the examination both notations will be accepted.

Old notation	IA	IIA	IIIB	IVB	VB	VIB	VIIB	VIII	IB	IIB	IIIA	IVA	VA	VIA	VIIA	0		
New notation	1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18

1. Amino acids: Peptides, structure and classification, proteins structure and their role in the growth of living beings.
2. Nucleic Acid: DNA and RNA – their structure. Unique nature. Importance in evolution and their characteristic features.
3. Carbohydrates and their metabolism, Blood - haemoglobin and respiration.
4. Vitamins and hormones
5. Simple idea of chemical evolution.
6. Natural polymers (any **five**) - structure, characteristics, uses. Synthetic polymers (any **five**) - method of preparation, structure, characteristics and uses.
7. Types of Dyes - methods of preparation, characteristics and uses.
8. Chemicals in medicines: antiseptics, antibiotics, antacids, etc. and their uses.
9. Preparation of soap, nail polish, boot polish, varnish, nail polish remover, shampoo and perfumes.
10. Chemicals and chemical processes in forensic studies.
11. Insecticides, pesticides and chemical fertilisers.
12. Ancient Indian medicines and medicinal plants.
13. Organic Chemistry in Nutrition, Food Science and Biotechnology.
14. Effect of Green House Gases.
15. How Plastics have changed the world, both socially and economically.

Practical File – 5 Marks

The Visiting Examiner is required to assess students on the basis of the Chemistry Practical file maintained by them during the academic year.

MATHEMATICS (860)

CLASS XII

*There will be **two** papers in the subject:*

Paper I : Theory (3 hours)80 marks

Paper II: Project Work20 marks

PAPER I (THEORY) – 80 Marks

*The syllabus is divided into **three** sections A, B and C.*

*Section A is compulsory for all candidates. Candidates will have a choice of attempting questions from **EITHER** Section B **OR** Section C.*

DISTRIBUTION OF MARKS FOR THE THEORY PAPER

S.No.	UNIT	TOTAL WEIGHTAGE
SECTION A: 65 MARKS		
1.	Relations and Functions	10 Marks
2.	Algebra	10 Marks
3.	Calculus	32 Marks
4.	Probability	13 Marks
SECTION B: 15 MARKS		
5.	Vectors	5 Marks
6.	Three - Dimensional Geometry	6 Marks
7.	Applications of Integrals	4 Marks
OR SECTION C: 15 MARKS		
8.	Application of Calculus	5 Marks
9.	Linear Regression	6 Marks
10.	Linear Programming	4 Marks
TOTAL		80 Marks

SECTION A

1. Relations and Functions

(i) Types of relations: reflexive, symmetric, transitive and equivalence relations. One to one and onto functions, inverse of a function.

- *Relations as:*

- *Relation on a set A*
- *Identity relation, empty relation, universal relation.*
- *Types of Relations: reflexive, symmetric, transitive and equivalence relation.*

- *Functions:*

- *As special relations, concept of writing "y is a function of x" as $y = f(x)$.*
- *Types: one to one, many to one, into, onto.*
- *Real Valued function.*
- *Domain and range of a function.*
- *Conditions of invertibility.*
- *Invertible functions (algebraic functions only).*

(ii) Inverse Trigonometric Functions

Definition, domain, range, principal value branch. Elementary properties of inverse trigonometric functions.

- *Principal values.*
- $\sin^{-1}x, \cos^{-1}x, \tan^{-1}x$ etc
- $\sin^{-1}x = \cos^{-1}\sqrt{1-x^2} = \tan^{-1}\frac{x}{\sqrt{1-x^2}}$.
- $\sin^{-1}x = \operatorname{cosec}^{-1}\frac{1}{x}; \sin^{-1}x + \cos^{-1}x = \frac{\pi}{2}$ and similar relations for $\cot^{-1}x, \tan^{-1}x$, etc.

$$\sin^{-1}x \pm \sin^{-1}y = \sin^{-1}\left(x\sqrt{1-y^2} \pm y\sqrt{1-x^2}\right)$$

$$\cos^{-1}x \pm \cos^{-1}y = \cos^{-1}\left(xy \mp \sqrt{1-y^2}\sqrt{1-x^2}\right)$$

$$\text{similarly } \tan^{-1}x + \tan^{-1}y = \tan^{-1}\frac{x+y}{1-xy}, xy < 1$$

$$\tan^{-1}x - \tan^{-1}y = \tan^{-1}\frac{x-y}{1+xy}, xy > -1$$

- *Formulae for $2\sin^{-1}x, 2\cos^{-1}x, 2\tan^{-1}x, 3\tan^{-1}x$ etc. and application of these formulae.*

2. Algebra

Matrices and Determinants

(i) Matrices

Concept, notation, order, equality, types of matrices, zero and identity matrix, transpose of a matrix, symmetric and skew symmetric matrices. Operation on matrices: Addition and multiplication and multiplication with a scalar. Simple properties of addition, multiplication and scalar multiplication. Non-commutativity of multiplication of matrices and existence of non-zero matrices whose product is the zero matrix (restrict to square matrices of order upto 3). Invertible matrices and proof of the uniqueness of inverse, if it exists (here all matrices will have real entries).

(ii) Determinants

Determinant of a square matrix (up to 3×3 matrices), properties of determinants, minors, co-factors and applications of determinants in finding the area of a triangle. Adjoint and inverse of a square matrix. Consistency, inconsistency and number of solutions of system of linear equations by examples, solving system of linear equations in two or three variables (having unique solution) using inverse of a matrix.

- *Types of matrices ($m \times n$; $m, n \leq 3$), order; Identity matrix, Diagonal matrix.*
- *Symmetric, Skew symmetric.*
- *Operation – addition, subtraction, multiplication of a matrix with scalar,*

multiplication of two matrices (the compatibility).

E.g. $\begin{bmatrix} 1 & 1 \\ 0 & 2 \\ 1 & 1 \end{bmatrix} \begin{bmatrix} 1 & 2 \\ 2 & 2 \end{bmatrix} = AB$ (say) but BA is not possible.

- Singular and non-singular matrices.
- Existence of two non-zero matrices whose product is a zero matrix.
- Inverse ($2 \times 2, 3 \times 3$) $A^{-1} = \frac{\text{Adj}A}{|A|}$

- Martin's Rule (i.e. using matrices)

$$a_1x + b_1y + c_1z = d_1$$

$$a_2x + b_2y + c_2z = d_2$$

$$a_3x + b_3y + c_3z = d_3$$

$$A = \begin{bmatrix} a_1 & b_1 & c_1 \\ a_2 & b_2 & c_2 \\ a_3 & b_3 & c_3 \end{bmatrix} \quad B = \begin{bmatrix} d_1 \\ d_2 \\ d_3 \end{bmatrix} \quad X = \begin{bmatrix} x \\ y \\ z \end{bmatrix}$$

$$AX = B \Rightarrow X = A^{-1}B$$

Problems based on above.

NOTE: The conditions for consistency of equations in two and three variables, using matrices, are to be covered.

- Determinants

- Order.
- Minors.
- Cofactors.
- Expansion.
- Applications of determinants in finding the area of triangle and collinearity.
- Properties of determinants. Problems based on properties of determinants.

3. Calculus

- (i) Continuity, Differentiability and Differentiation. Continuity and differentiability, derivative of composite functions, chain rule, derivatives of inverse trigonometric functions, derivative of implicit functions. Concept of exponential and logarithmic functions.

Derivatives of logarithmic and exponential functions. Logarithmic differentiation, derivative of functions expressed in parametric forms. Second order derivatives.

- Continuity

- Continuity of a function at a point $x = a$.
- Continuity of a function in an interval.
- Algebra of continuous function.
- Removable discontinuity.

- Differentiation

- Concept of continuity and differentiability of $|x|$, $[x]$, etc.
- Derivatives of trigonometric functions.
- Derivatives of exponential functions.
- Derivatives of logarithmic functions.
- Derivatives of inverse trigonometric functions - differentiation by means of substitution.
- Derivatives of implicit functions and chain rule.
- Derivatives of Parametric functions.
- Differentiation of a function with respect to another function e.g. differentiation of $\sin x^3$ with respect to x^3 .
- Logarithmic Differentiation - Finding dy/dx when $y = x^{x^x}$.
- Successive differentiation up to 2nd order.

NOTE: Derivatives of composite functions using chain rule.

(ii) Applications of Derivatives

Applications of derivatives: rate of change of bodies, increasing/decreasing functions, tangents and normals, maxima and minima (first derivative test motivated geometrically and second derivative test given as a provable tool). Simple problems (that illustrate basic principles and understanding of the subject as well as real-life situations).

- *Equation of Tangent and Normal*
- *Rate measure.*
- *Increasing and decreasing functions.*
- *Maxima and minima.*
 - *Stationary/turning points.*
 - *Absolute maxima/minima*
 - *local maxima/minima*
 - *First derivatives test and second derivatives test*
 - *Application problems based on maxima and minima.*

(iii) Integrals

Integration as inverse process of differentiation. Integration of a variety of functions by substitution, by partial fractions and by parts, Evaluation of simple integrals of the following types and problems based on them.

Fundamental Theorem of Calculus (without proof). Basic properties of definite integrals and evaluation of definite integrals.

- *Indefinite integral*
 - *Integration as the inverse of differentiation.*
 - *Anti-derivatives of polynomials and functions $(ax+b)^n$, $\sin x$, $\cos x$, $\sec^2 x$, $\operatorname{cosec}^2 x$ etc.*
 - *Integrals of the type $\sin^2 x$, $\sin^3 x$, $\sin^4 x$, $\cos^2 x$, $\cos^3 x$, $\cos^4 x$.*
 - *Integration of $1/x$, e^x .*
 - *Integration by substitution.*

- *Integrals of the type $f'(x)[f(x)]^n$, $\frac{f'(x)}{f(x)}$.*

- *Integration of $\tan x$, $\cot x$, $\sec x$, $\operatorname{cosec} x$.*

- *Integration by parts.*

- *Integration using partial fractions.*

Expressions of the form $\frac{f(x)}{g(x)}$ when degree of $f(x) < \text{degree of } g(x)$

E.g.
$$\frac{x+2}{(x-3)(x+1)} = \frac{A}{x-3} + \frac{B}{x+1}$$

$$\frac{x+2}{(x-2)(x-1)^2} = \frac{A}{x-1} + \frac{B}{(x-1)^2} + \frac{C}{x-2}$$

$$\frac{x+1}{(x^2+3)(x-1)} = \frac{Ax+B}{x^2+3} + \frac{C}{x-1}$$

When degree of $f(x) \geq \text{degree of } g(x)$,

e.g.
$$\frac{x^2+1}{x^2+3x+2} = 1 - \left(\frac{3x+1}{x^2+3x+2} \right)$$

- *Integrals of the type:*

$$\int \frac{dx}{x^2 \pm a^2}, \int \frac{dx}{\sqrt{x^2 \pm a^2}}, \int \frac{px+q}{ax^2+bx+c} dx, \int \frac{px+q}{\sqrt{ax^2+bx+c}} dx$$

and $\int \sqrt{a^2 \pm x^2} dx, \int \sqrt{x^2 - a^2} dx,$

$$\int \sqrt{ax^2+bx+c} dx, \int (px+q)\sqrt{ax^2+bx+c} dx,$$

integrations reducible to the above forms.

$$\int \frac{dx}{a \cos x + b \sin x},$$

$$\int \frac{dx}{a+b \cos x}, \int \frac{dx}{a+b \sin x} \int \frac{dx}{a \cos x + b \sin x + c},$$

$$\int \frac{(a \cos x + b \sin x) dx}{c \cos x + d \sin x},$$

$$\int \frac{dx}{a \cos^2 x + b \sin^2 x + c}$$

$$\int \frac{1 \pm x^2}{1+x^4} dx,$$

$$\int \frac{dx}{1+x^4}, \int \sqrt{\tan x} dx, \int \sqrt{\cot x} dx \text{ etc.}$$

- *Definite Integral*

- *Fundamental theorem of calculus (without proof)*
- *Properties of definite integrals.*
- *Problems based on the following properties of definite integrals are to be covered.*

$$\int_a^b f(x)dx = \int_a^b f(t)dt$$

$$\int_a^b f(x)dx = - \int_b^a f(x)dx$$

$$\int_a^b f(x)dx = \int_a^c f(x)dx + \int_c^b f(x)dx$$

where $a < c < b$

$$\int_a^b f(x)dx = \int_a^b f(a+b-x)dx$$

$$\int_0^a f(x)dx = \int_0^a f(a-x)dx$$

$$\int_0^{2a} f(x)dx = \begin{cases} 2 \int_0^a f(x)dx, & \text{if } f(2a-x) = f(x) \\ 0, & \text{if } f(2a-x) = -f(x) \end{cases}$$

$$\int_{-a}^a f(x)dx = \begin{cases} 2 \int_0^a f(x)dx, & \text{if } f \text{ is an even function} \\ 0, & \text{if } f \text{ is an odd function} \end{cases}$$

(iv) *Differential Equations*

Definition, order and degree, general and particular solutions of a differential equation. Solution of differential equations by method of separation of variables solutions of homogeneous differential equations of first order and first degree. Solutions of linear differential equation of the type: $\frac{dy}{dx} + py = q$, where p and q are

functions of x or constants. $\frac{dx}{dy} + px = q$,

where p and q are functions of y or constants.

- *Differential equations, order and degree.*
- *Formation of differential equation by eliminating arbitrary constant(s).*
- *Solution of differential equations.*
- *Variable separable.*
- *Homogeneous equations.*

- *Linear form $\frac{dy}{dx} + Py = Q$ where P and Q are functions of x only. Similarly, for dx/dy .*

NOTE 1: Equations reducible to variable separable type are included.

NOTE 2: The second order differential equations are excluded.

4. Probability

Conditional probability, multiplication theorem on probability, independent events, total probability, Bayes' theorem, Random variable and its probability distribution, mean of random variable.

- *Independent and dependent events conditional events.*
- *Laws of Probability, addition theorem, multiplication theorem, conditional probability.*
- *Theorem of Total Probability.*
- *Baye's theorem.*
- *Theoretical probability distribution, probability distribution function; mean of random variable.*

SECTION B

5. Vectors

Vectors and scalars, magnitude and direction of a vector. Direction cosines and direction ratios of a vector. Types of vectors (equal, unit, zero, parallel and collinear vectors), position vector of a point, negative of a vector, components of a vector, addition of vectors, multiplication of a vector by a scalar, position vector of a point dividing a line segment in a given ratio. Definition, Geometrical Interpretation, properties and application of scalar (dot) product of vectors, vector (cross) product of vectors.

- As directed line segments.
- Magnitude and direction of a vector.
- Types: equal vectors, unit vectors, zero vector.
- Position vector.
- Components of a vector.
- Vectors in two and three dimensions.
- $\hat{i}, \hat{j}, \hat{k}$ as unit vectors along the x, y and the z axes; expressing a vector in terms of the unit vectors.
- Operations: Sum and Difference of vectors; scalar multiplication of a vector.
- Section formula.
- Scalar (dot) product of vectors and its geometrical significance.
- Cross product - its properties - area of a triangle, area of parallelogram, collinear vectors.

NOTE: Proofs of geometrical theorems by using Vector algebra are excluded.

6. Three - dimensional Geometry

Direction cosines and direction ratios of a line joining two points. Cartesian equation and vector equation of a line, coplanar and skew lines, shortest distance between two lines. Cartesian and vector equation of a plane. Angle between (i) two lines, (ii) two planes, (iii) a line and a plane. Distance of a point from a plane.

- Equation of x-axis, y-axis, z axis and lines parallel to them.
- Equation of xy - plane, yz - plane, zx - plane.
- Direction cosines, direction ratios.
- Angle between two lines in terms of direction cosines /direction ratios.
- Condition for lines to be perpendicular/ parallel.
- Lines
 - Cartesian and vector equations of a line through one and two points.
 - Coplanar and skew lines.
 - Conditions for intersection of two lines.

- Distance of a point from a line.
- Shortest distance between two lines.
- Planes
 - Cartesian and vector equation of a plane.
 - Direction ratios of the normal to the plane.
 - One point form.
 - Normal form.
 - Intercept form.
 - Distance of a point from a plane.
 - Intersection of the line and plane.
 - Angle between two planes, a line and a plane.

7. Application of Integrals

Application in finding the area bounded by simple curves and coordinate axes. Area enclosed between two curves.

- Application of definite integrals - area bounded by curves, lines and coordinate axes is required to be covered.
- Simple curves: lines, circles/ parabolas/ ellipses, polynomial functions, modulus function.

SECTION C

8. Application of Calculus

Application of Calculus in Commerce and Economics in the following:

- Cost function,
- average cost,
- marginal cost and its interpretation
- demand function,
- revenue function,
- marginal revenue function and its interpretation,
- Profit function and breakeven point.
- Rough sketching of the following curves: AR, MR, R, C, AC, MC and their mathematical interpretation using the concept of maxima & minima and increasing-decreasing functions.

Self-explanatory

NOTE: Application involving differentiation, increasing and decreasing function and maxima and minima to be covered.

9. Linear Regression

- Lines of regression of x on y and y on x.
- Scatter diagrams
- The method of least squares.
- Lines of best fit.
- Regression coefficient of x on y and y on x.
- $b_{xy} \times b_{yx} = r^2, 0 \leq b_{xy} \times b_{yx} \leq 1$
- Identification of regression equations
- Properties of regression lines.
- Estimation of the value of one variable using the value of other variable from appropriate line of regression.

Self-explanatory

10. Linear Programming

Introduction, related terminology such as constraints, objective function, optimization, different types of linear programming (L.P.) problems, mathematical formulation of L.P. problems, graphical method of solution for problems in two variables, feasible and infeasible regions (bounded and unbounded), feasible and infeasible solutions, optimal feasible solutions (up to three non-trivial constraints).

Introduction, definition of related terminology such as constraints, objective function, optimization, advantages of linear programming; limitations of linear programming; application areas of linear programming; different types of linear programming (L.P.) problems, mathematical formulation of L.P. problems, graphical method of solution for problems in two variables, feasible and infeasible regions, feasible and infeasible solutions, optimum feasible solution.

PAPER II – PROJECT WORK – 20 Marks

Candidates will be expected to have completed two projects, one from Section A and one from either Section B or Section C.

The project work will be assessed by the subject teacher and a Visiting Examiner appointed locally and approved by the Council.

Mark allocation for each Project [10 marks]:

Overall format	1 mark
Content	4 marks
Findings	2 marks
Viva-voce based on the Project	3 marks
Total	10 marks

List of suggested assignments for Project Work:

Section A

1. Using a graph, demonstrate a function which is one-one but not onto.
2. Using a graph demonstrate a function which is invertible.
3. Construct a composition table using a binary function addition/multiplication modulo upto 5 and verify the existence of the properties of binary operation.
4. Draw the graph of $y = \sin^{-1} x$ (or any other inverse trigonometric function), using the graph of $y = \sin x$ (or any other relevant trigonometric function). Demonstrate the concept of mirror line (about $y = x$) and find its domain and range.
5. Explore the principal value of the function $\sin^{-1} x$ (or any other inverse trigonometric function) using a unit circle.
6. Find the derivatives of a determinant of the order of 3×3 and verify the same by other methods.
7. Verify the consistency of the system of three linear equations of two variables and verify the same graphically. Give its geometrical interpretation.
8. For a dependent system (non-homogeneous) of three linear equations of three variables, identify infinite number of solutions.
9. For a given function, give the geometrical interpretation of Mean Value theorems. Explain the significance of closed and open intervals for continuity and differentiability properties of the theorems.
10. Explain the concepts of increasing and decreasing functions, using geometrical

significance of dy/dx . Illustrate with proper examples.

11. Explain the geometrical significance of point of inflexion with examples and illustrate it using graphs.
12. Explain and illustrate (with suitable examples) the concept of local maxima and local minima using graph.
13. Explain and illustrate (with suitable examples) the concept of absolute maxima and absolute minima using graph.
14. Illustrate the concept of definite integral $\int_a^b f(x) dx$, expressing as the limit of a sum and verify it by actual integration.
15. Demonstrate application of differential equations to solve a given problem (example, population increase or decrease, bacteria count in a culture, etc.).
16. Explain the conditional probability, the theorem of total probability and the concept of Bayes' theorem with suitable examples.
17. Explain the types of probability distributions and derive mean and variance of binomial probability distribution for a given function.

Section B

18. Using Vector algebra, find the area of a parallelogram/triangle. Also, derive the area analytically and verify the same.
19. Using Vector algebra, prove the formulae of properties of triangles (sine/cosine rule, etc.)
20. Using Vector algebra, prove the formulae of compound angles, e.g. $\sin(A + B) = \sin A \cos B + \sin B \cos A$, etc.
21. Describe the geometrical interpretation of scalar triple product and for a given data, find the scalar triple product.
22. Find the image of a line with respect to a given plane.
23. Find the distance of a point from a given plane measured parallel to a given line.
24. Find the distance of a point from a line measured parallel to a given plane.

25. Find the area bounded by a parabola and an oblique line.
26. Find the area bounded by a circle and an oblique line.
27. Find the area bounded by an ellipse and an oblique line.
28. Find the area bounded by a circle and a circle.
29. Find the area bounded by a parabola and a parabola.
30. Find the area bounded by a circle and a parabola.

(Any other pair of curves which are specified in the syllabus may also be taken.)

Section C

31. Draw a rough sketch of Cost (C), Average Cost (AC) and Marginal Cost (MC)

Or

Revenue (R), Average Revenue (AR) and Marginal Revenue (MR).

Give their mathematical interpretation using the concept of increasing - decreasing functions and maxima-minima.
32. For a given data, find regression equations by the method of least squares. Also find angles between regression lines.
33. Draw the scatter diagram for a given data. Use it to draw the lines of best fit and estimate the value of Y when X is given and vice-versa.
34. Using any suitable data, find the minimum cost by applying the concept of Transportation problem.
35. Using any suitable data, find the minimum cost and maximum nutritional value by applying the concept of Diet problem.
36. Using any suitable data, find the Optimum cost in the manufacturing problem by formulating a linear programming problem (LPP).

NOTE: No question paper for Project Work will be set by CISCE.

SAMPLE TABLE FOR PROJECT WORK

S. No.	Unique Identification Number (Unique ID) of the candidate	<u>PROJECT 1</u>					<u>PROJECT 2</u>					TOTAL MARKS (E + J)
		A	B	C	D	E	F	G	H	I	J	
		Teacher	Visiting Examiner	Average Marks $(A + B \div 2)$	Viva-Voce by Visiting Examiner	Total Marks $(C + D)$	Teacher	Visiting Examiner	Average Marks $(F + G \div 2)$	Viva-Voce by Visiting Examiner	Total Marks $(H + I)$	
		7 Marks*	7 Marks*	7 Marks	3 Marks	10 Marks	7 Marks*	7 Marks*	7 Marks	3 Marks	10 Marks	
1												
2												
3												
4												
5												
6												
7												
8												
9												
10												

*Breakup of 7 Marks to be awarded separately by the Teacher and the Visiting Examiner is as follows:

Overall Format	1 Mark	Name of Teacher: Signature:	Date
Content	4 Marks	Name of Visiting Examiner	
Findings	2 Marks	Signature:	Date

NOTE: VIVA-VOCE (3 Marks) for each Project is to be conducted only by the Visiting Examiner, and should be based on the Project only